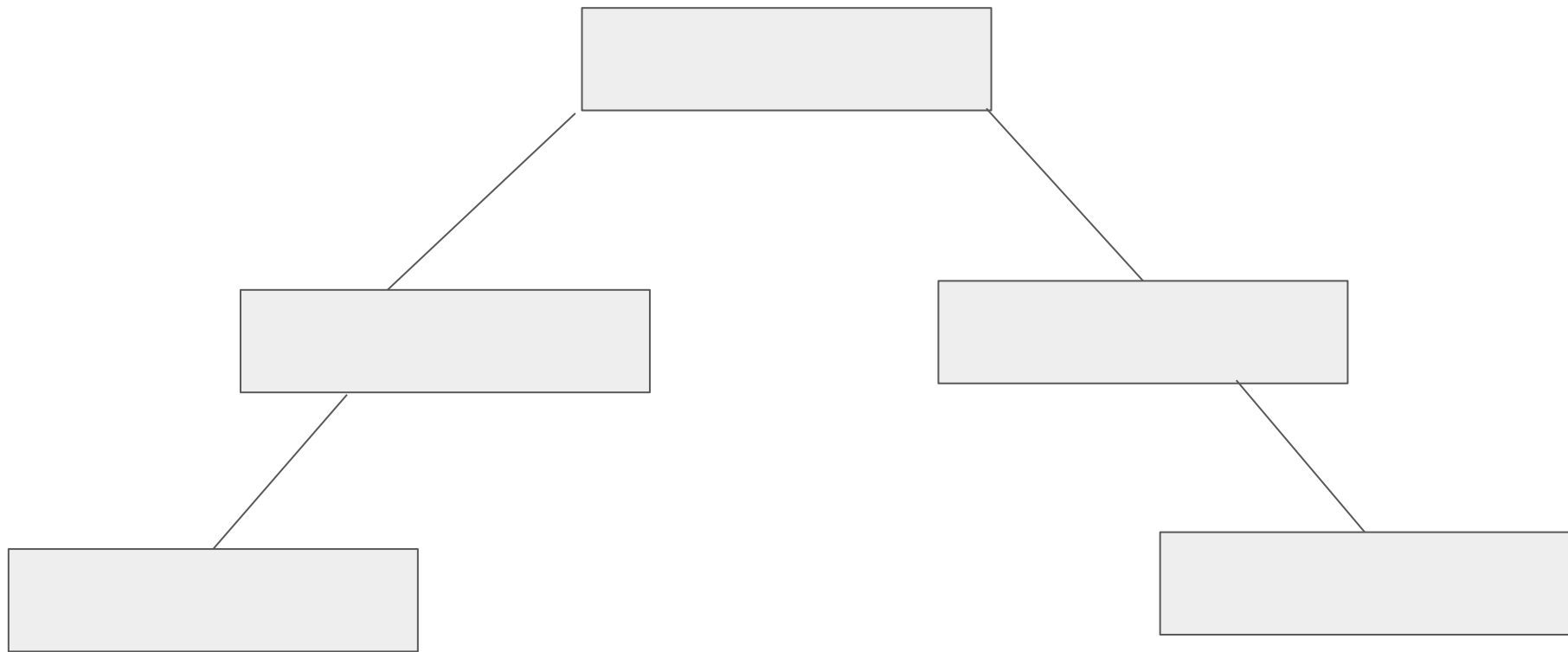




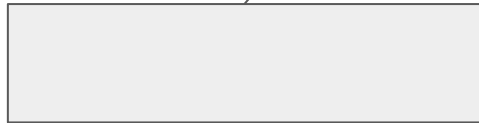
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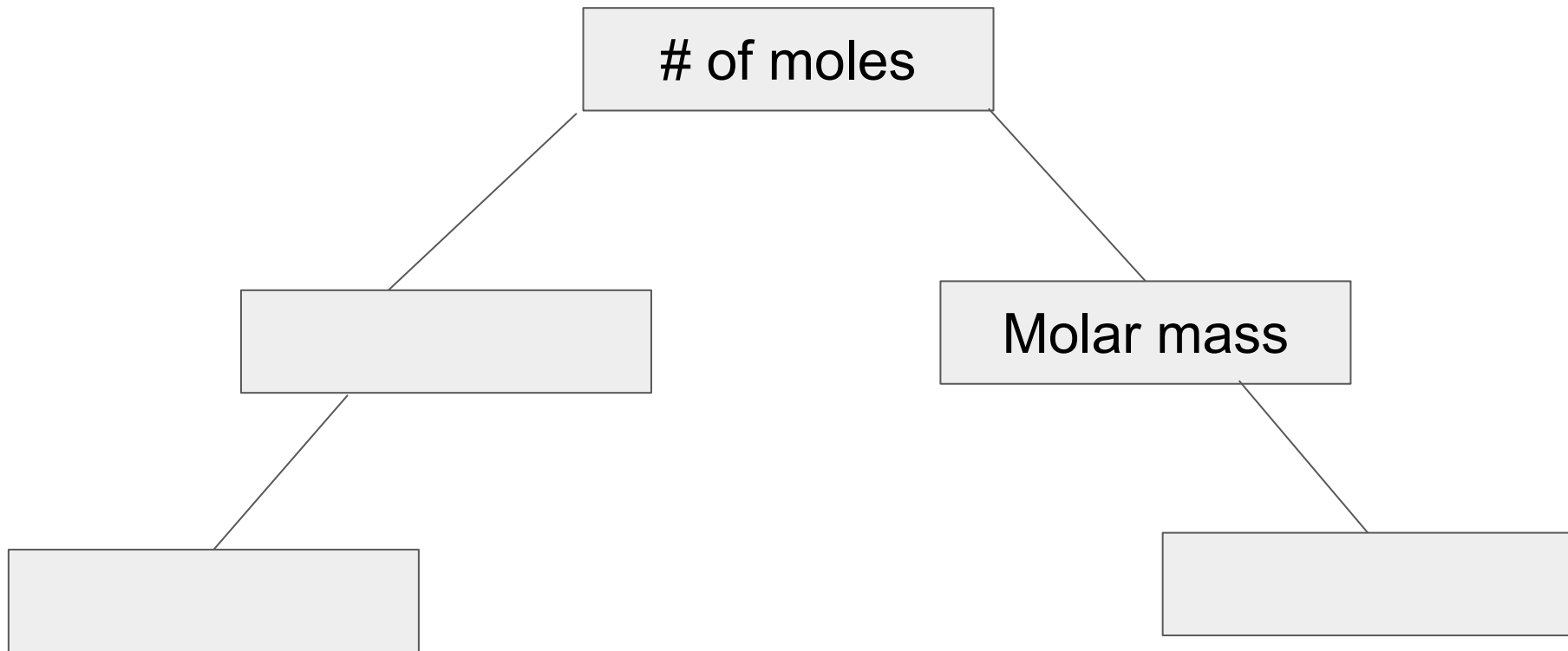
Molar Conversions the Easy Way

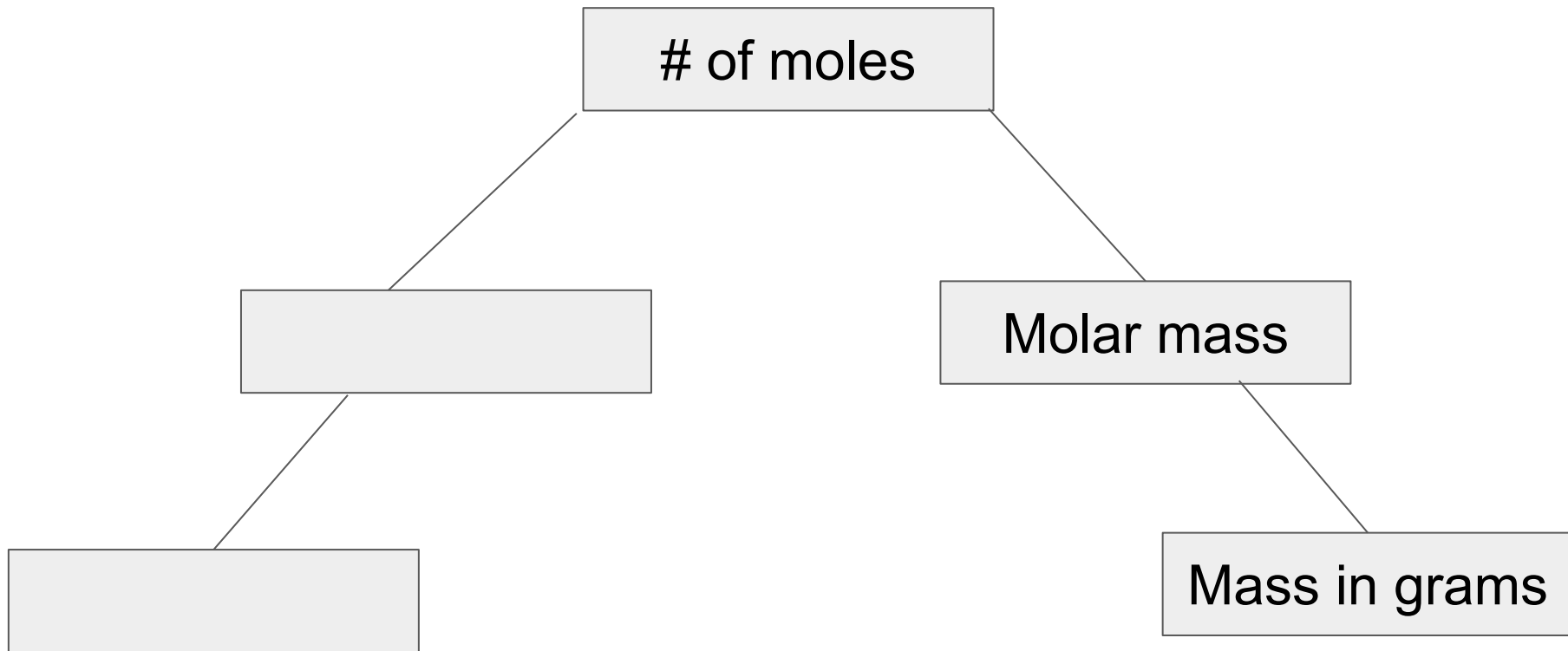


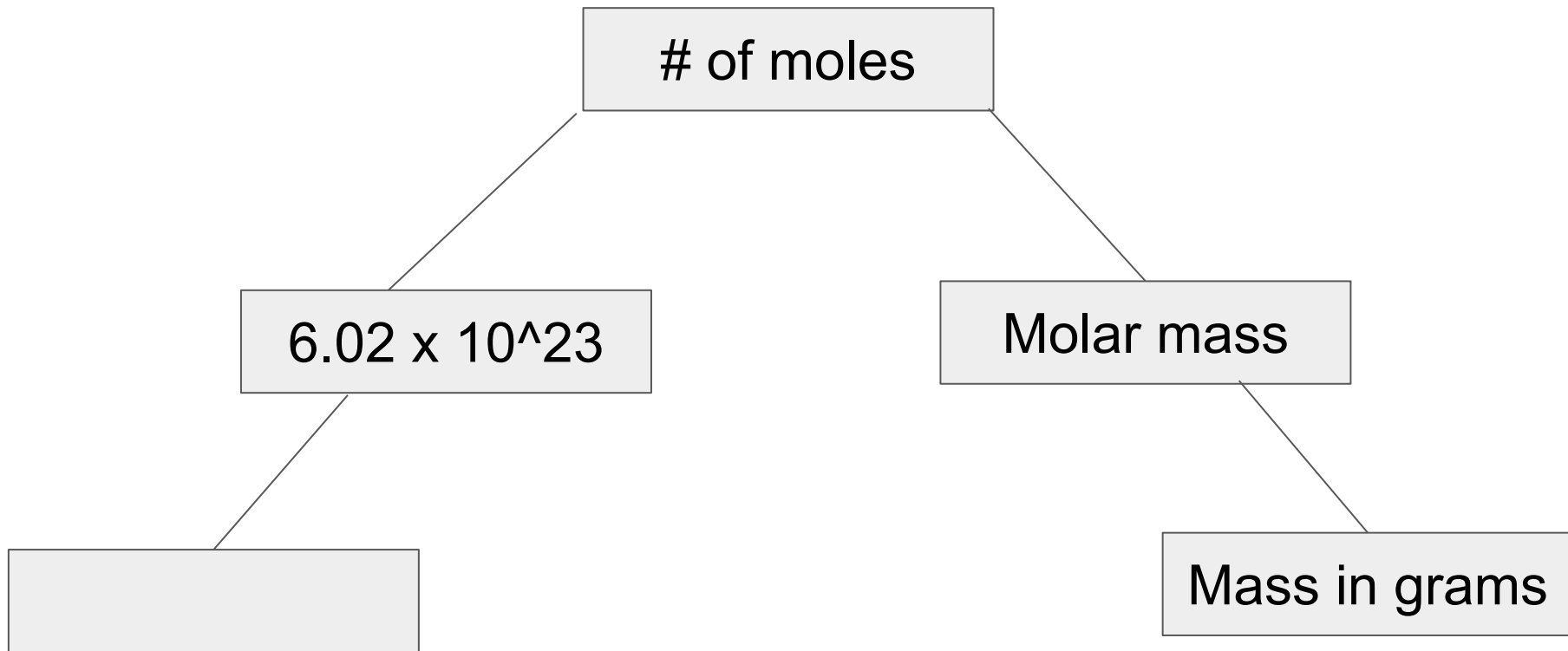
of moles

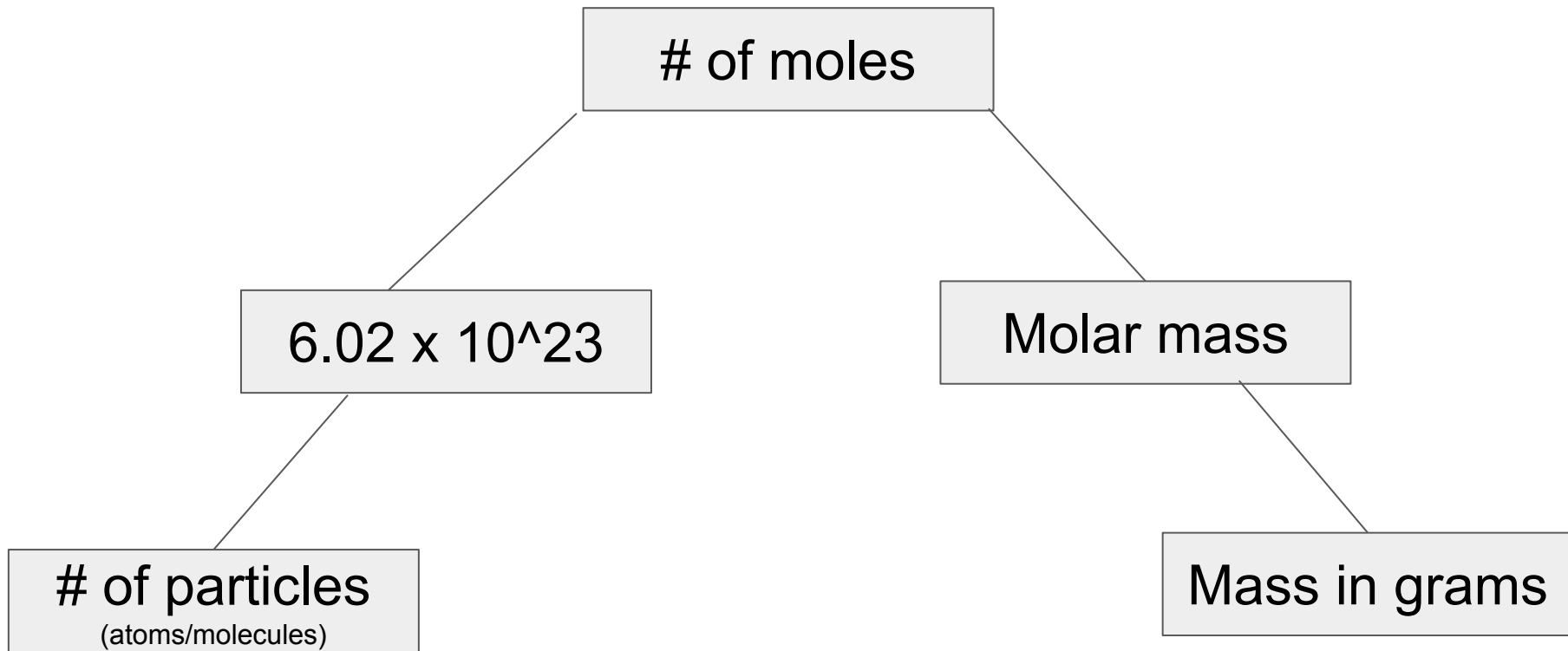
```
graph TD; A["# of moles"] --> B[" "]; A --> C[" "]; B --> D[" "]; C --> E[" "];
```





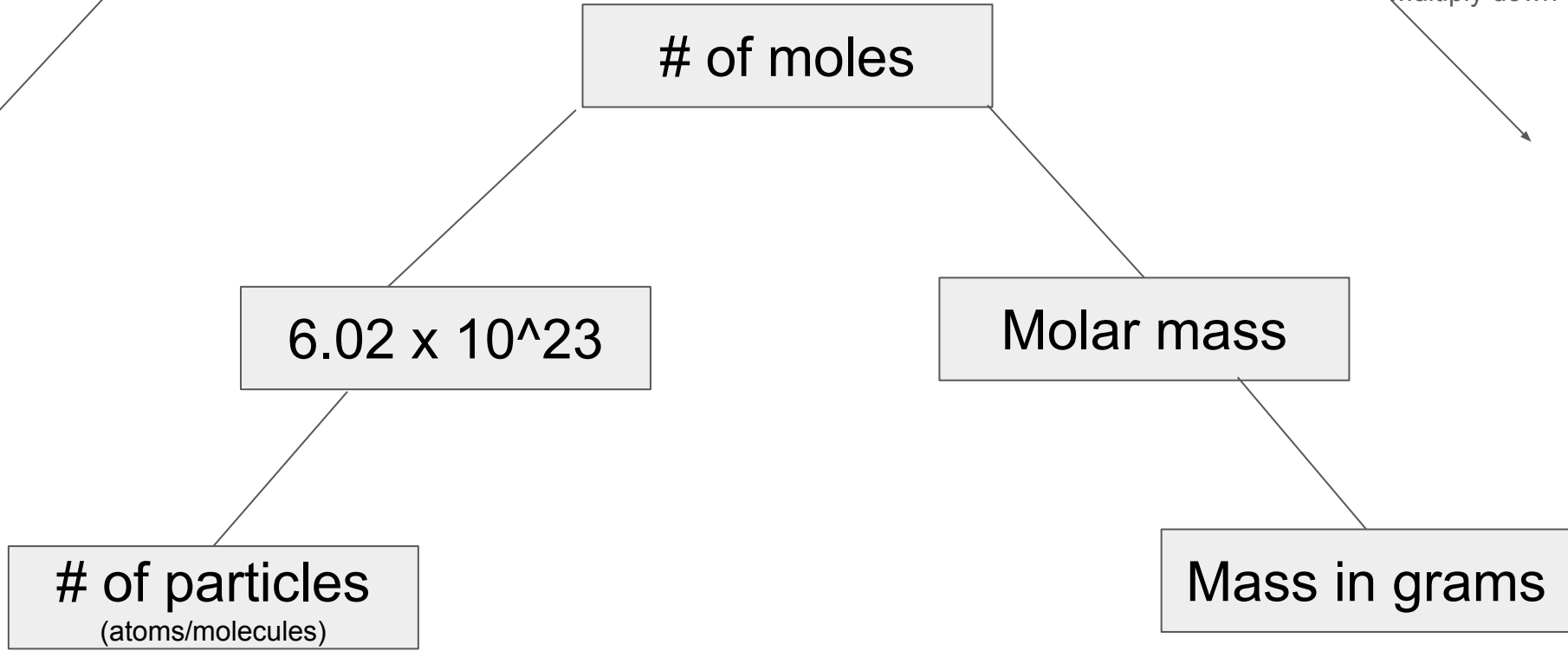


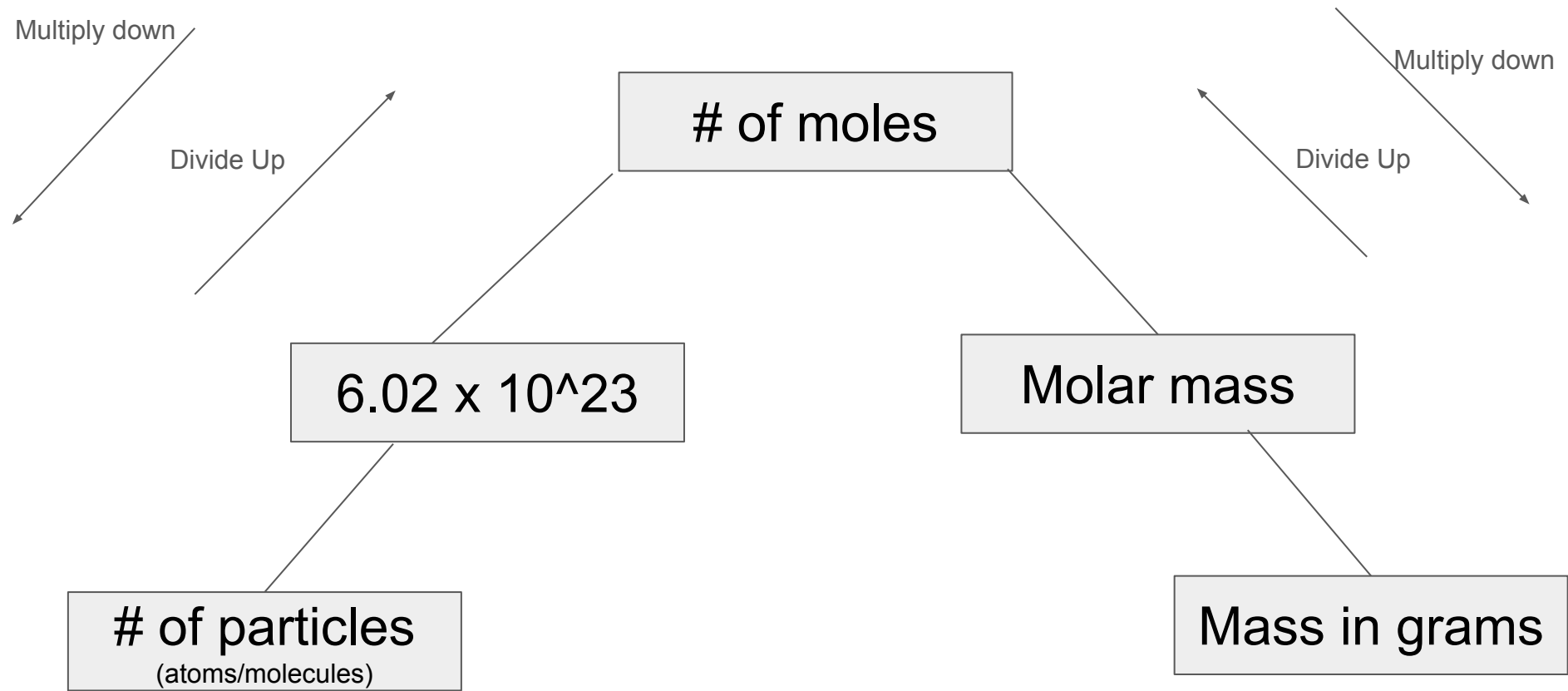


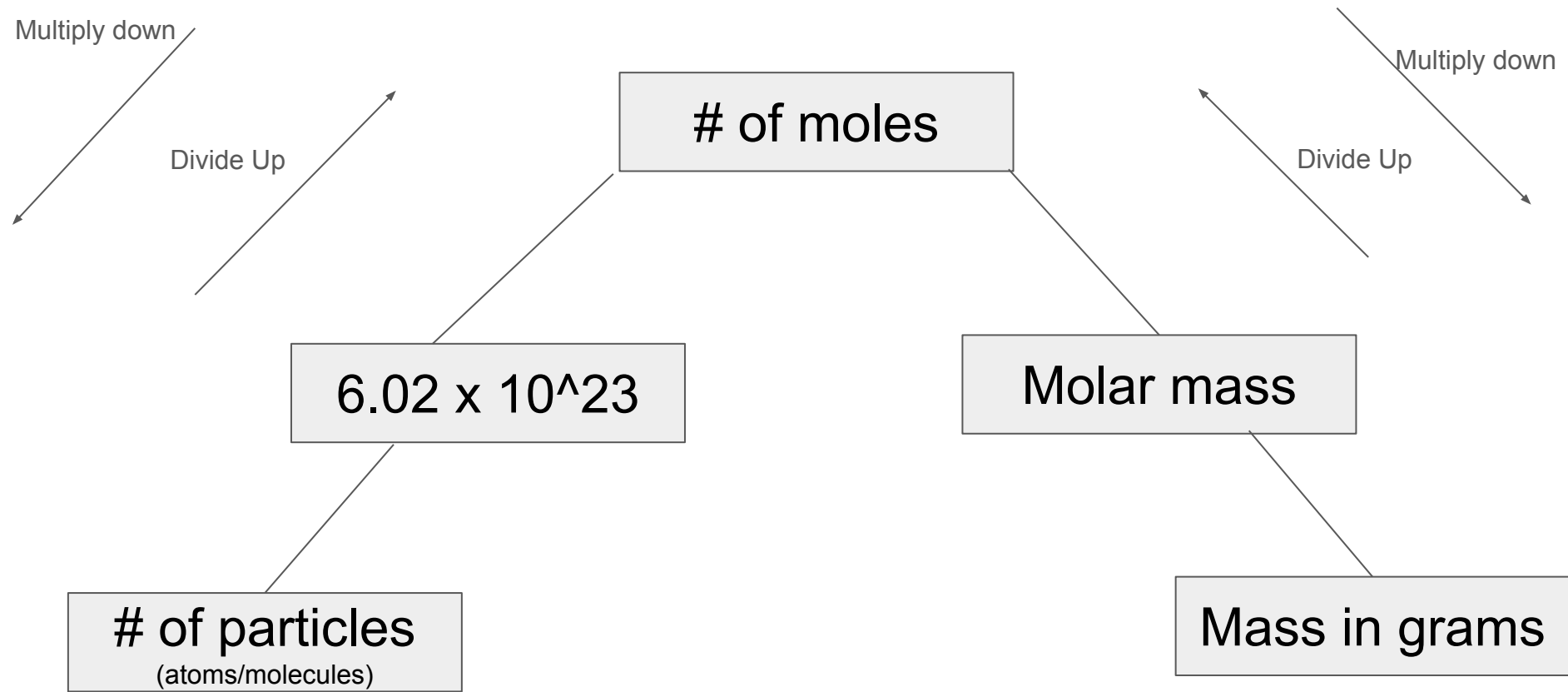


Multiply down
↙

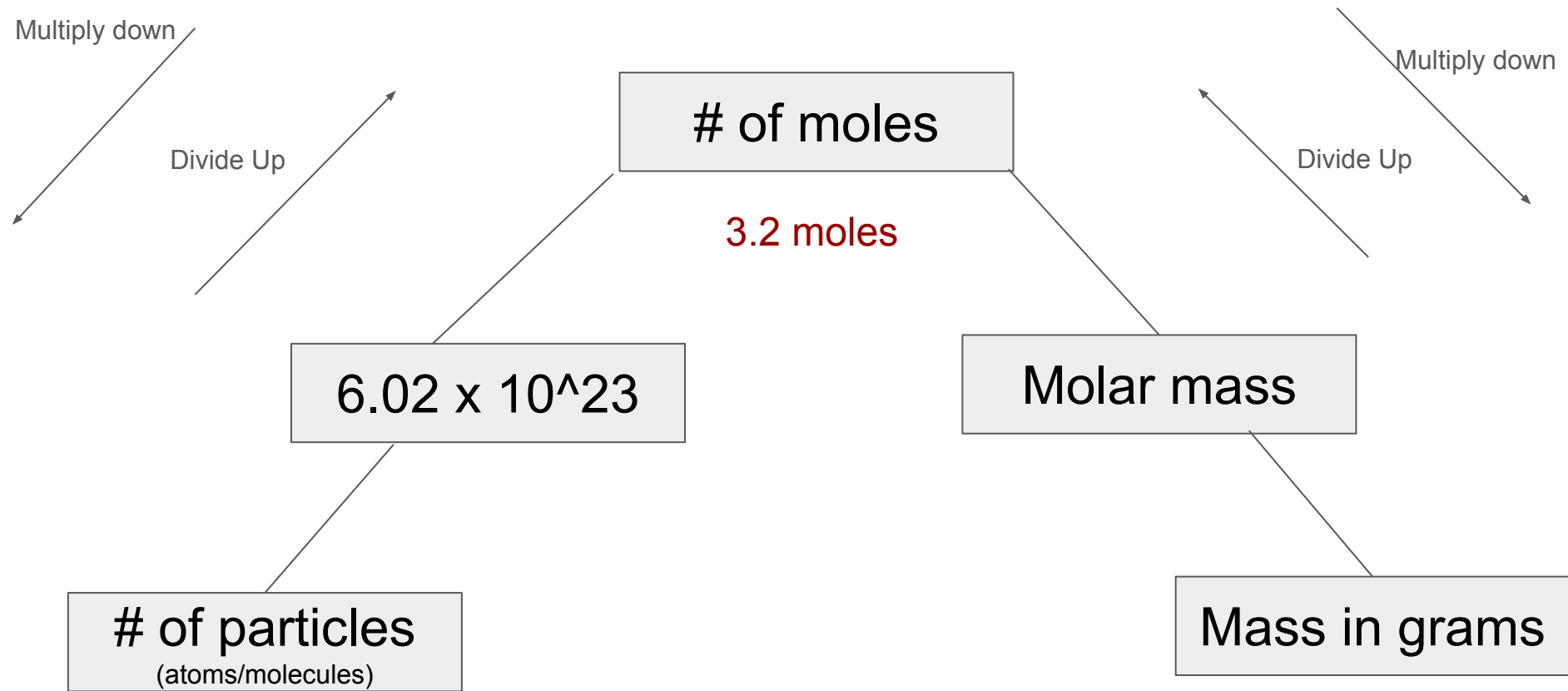
Multiply down
↘



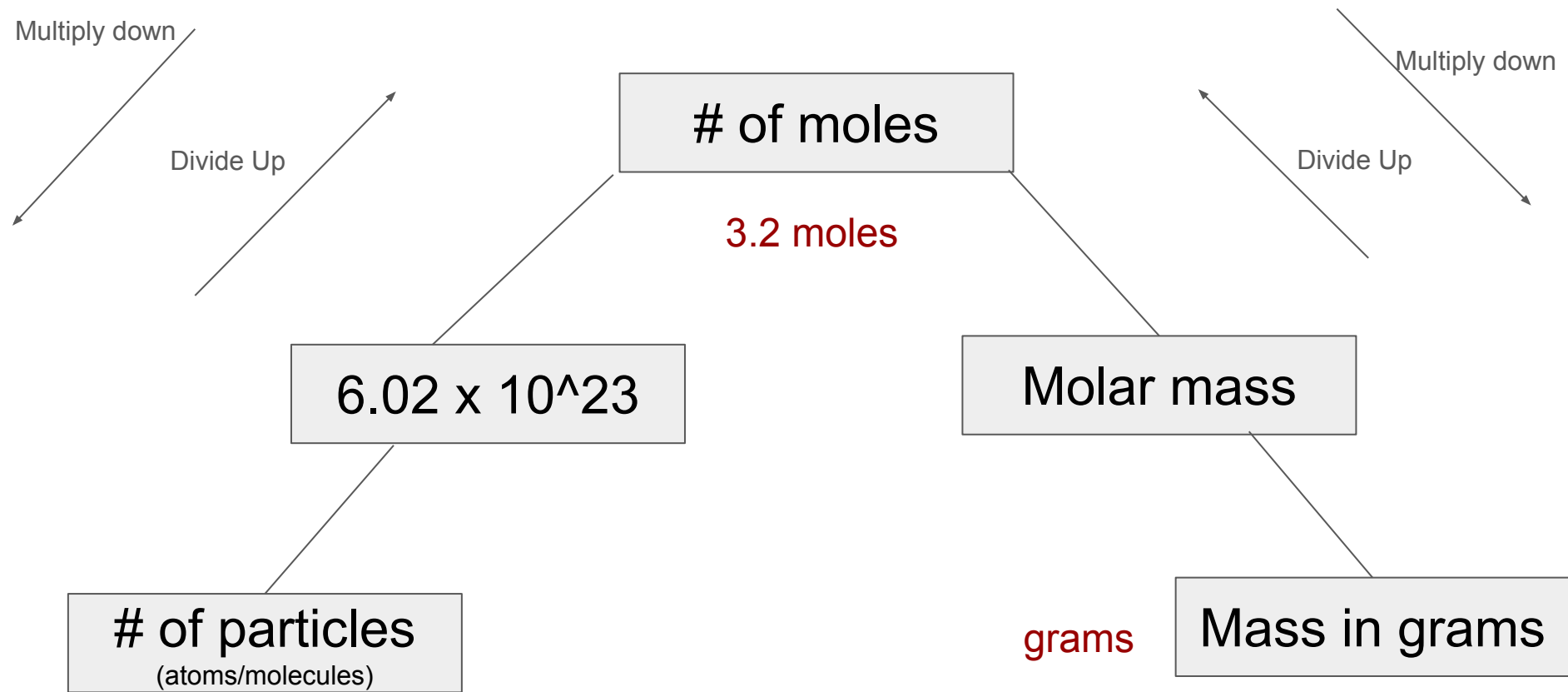




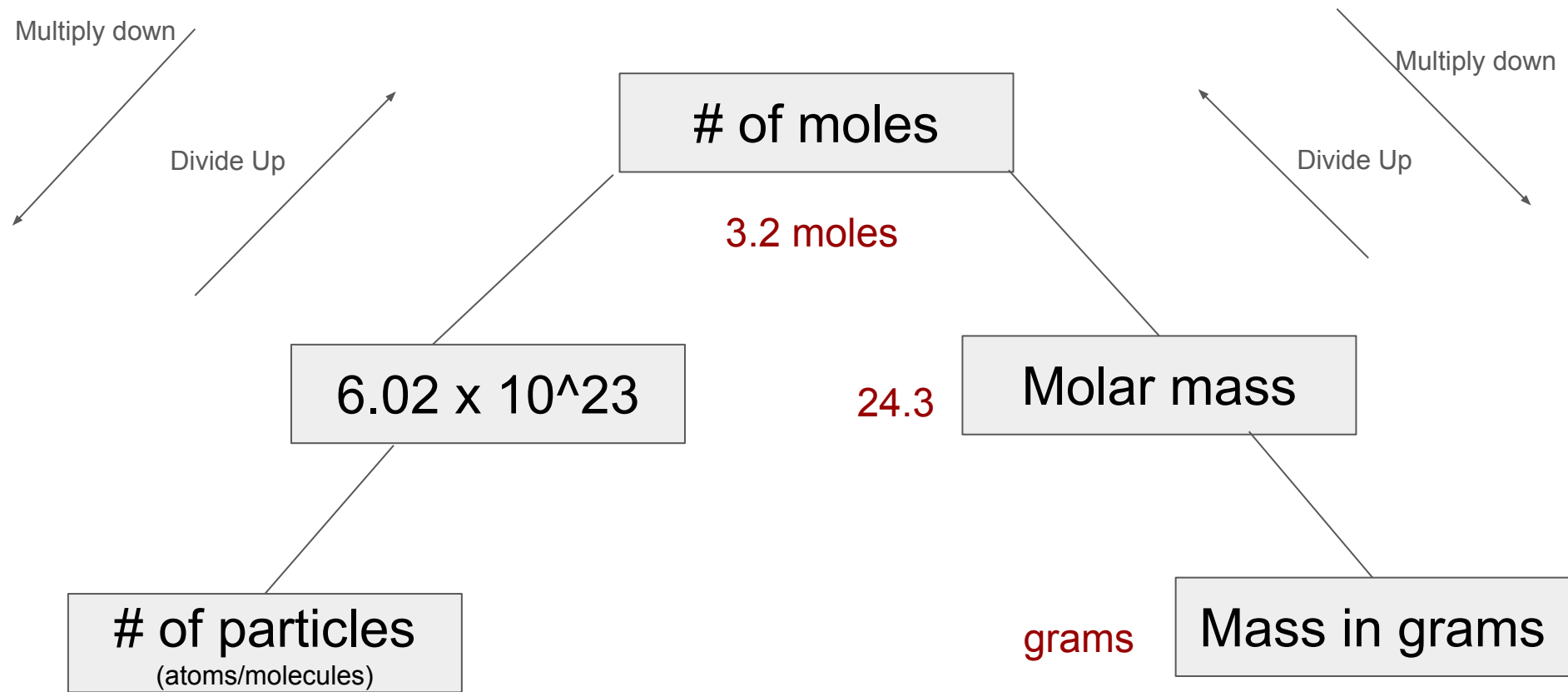
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



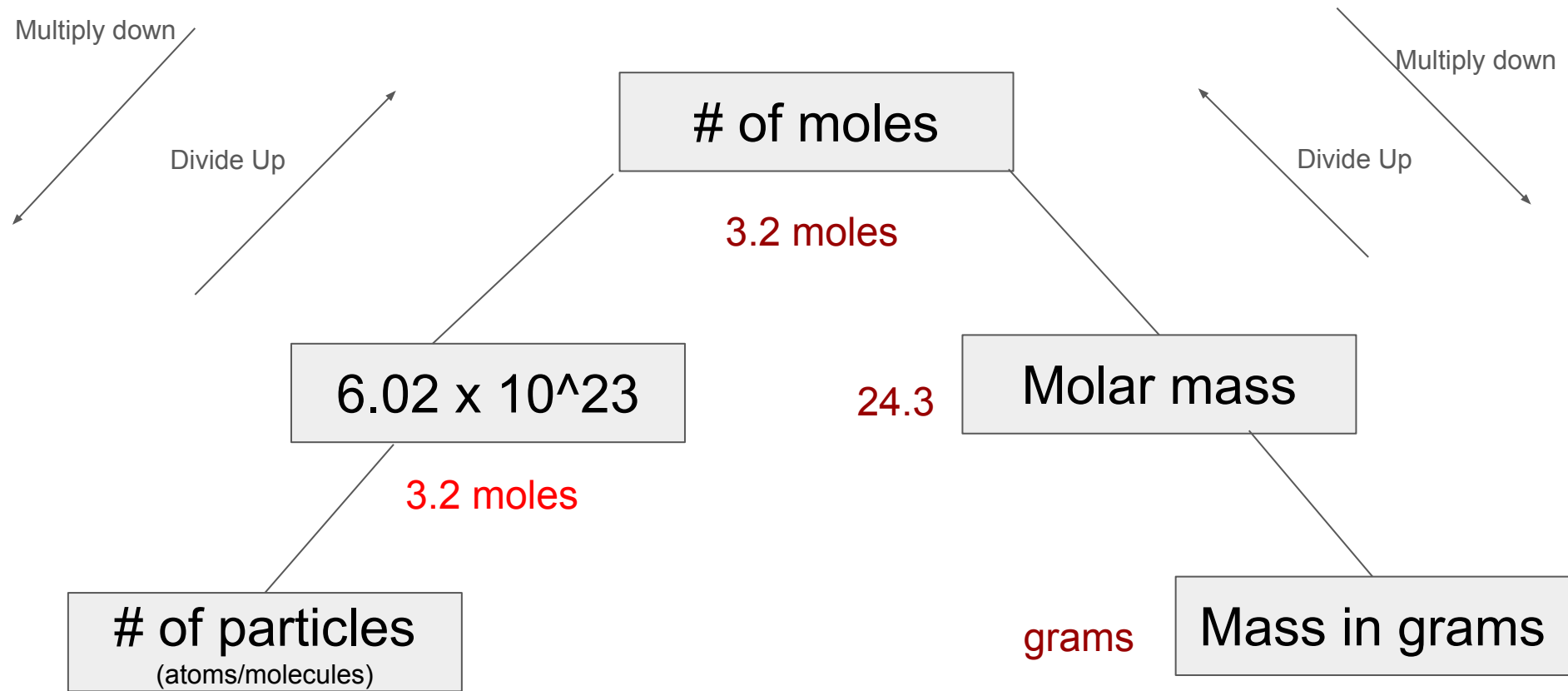
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



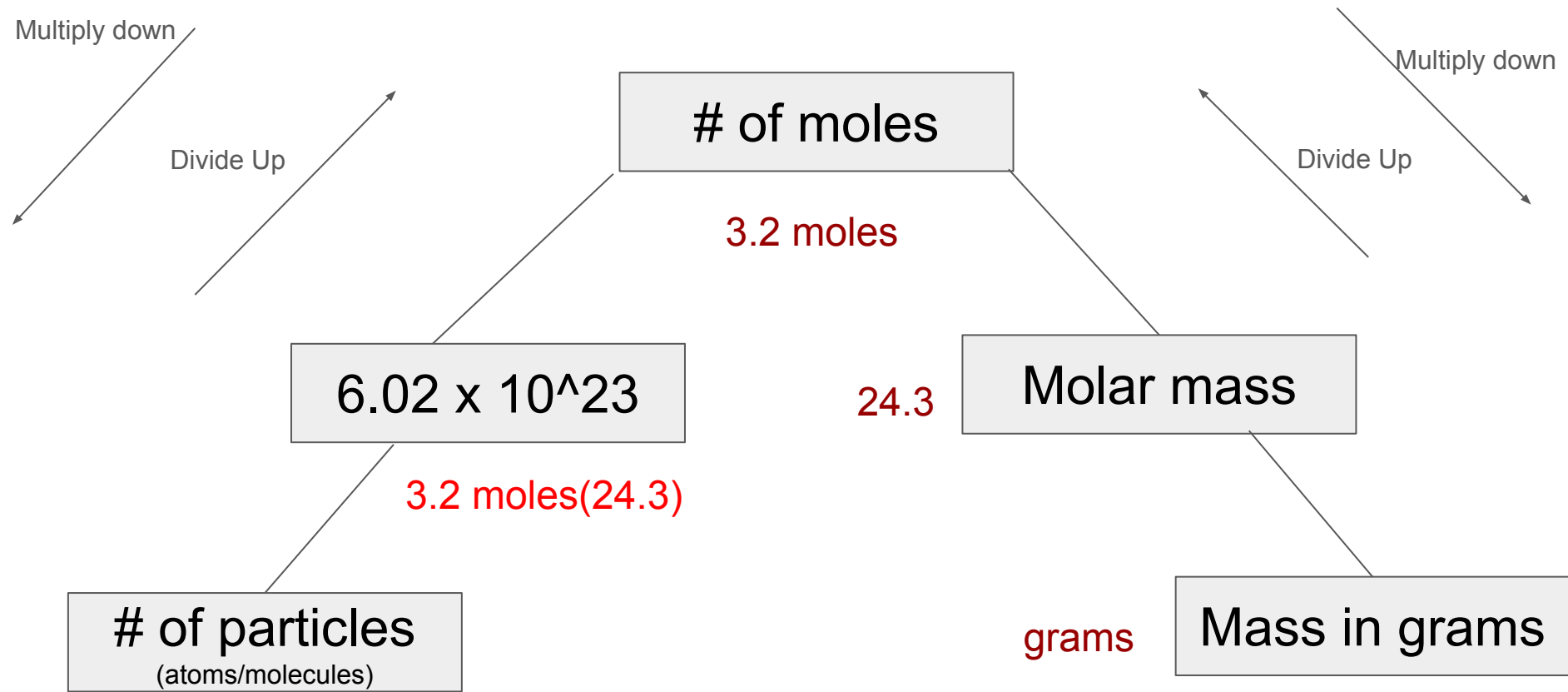
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



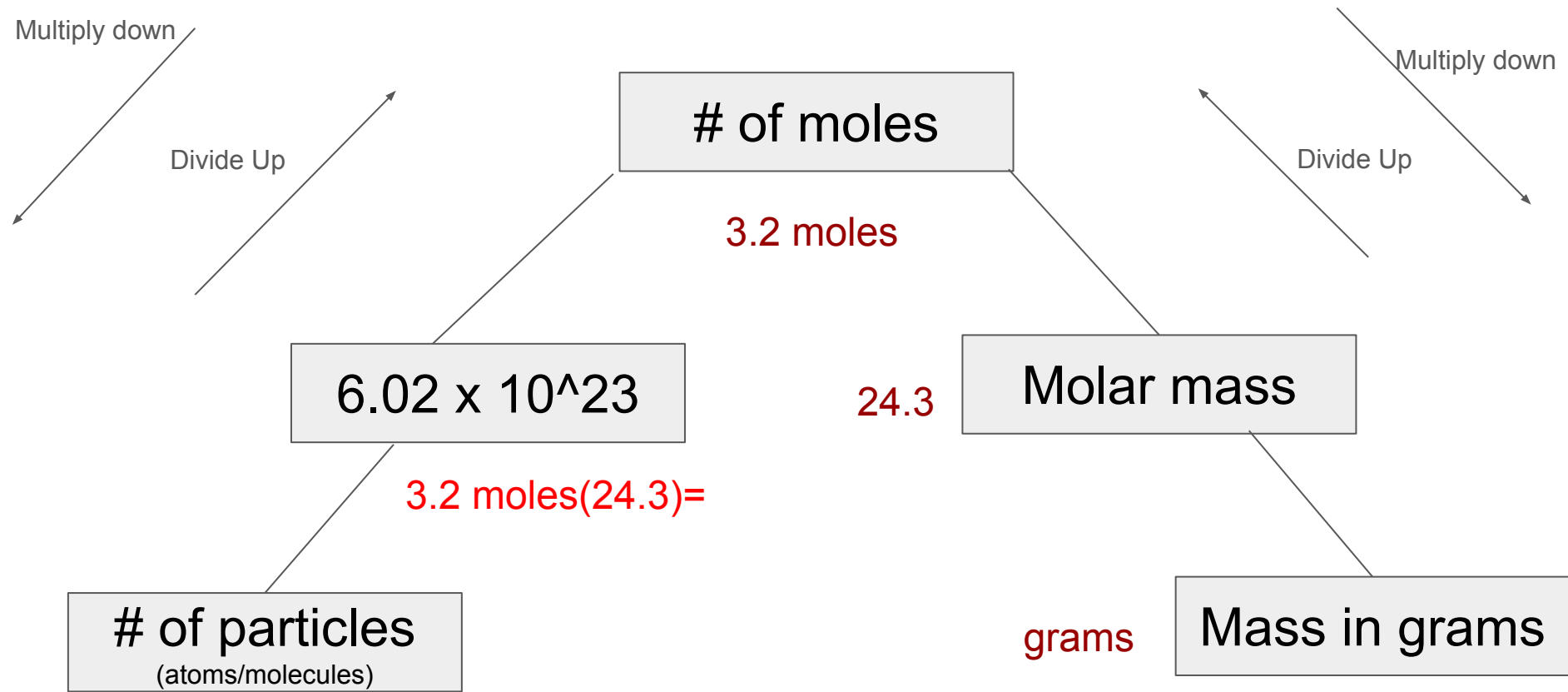
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



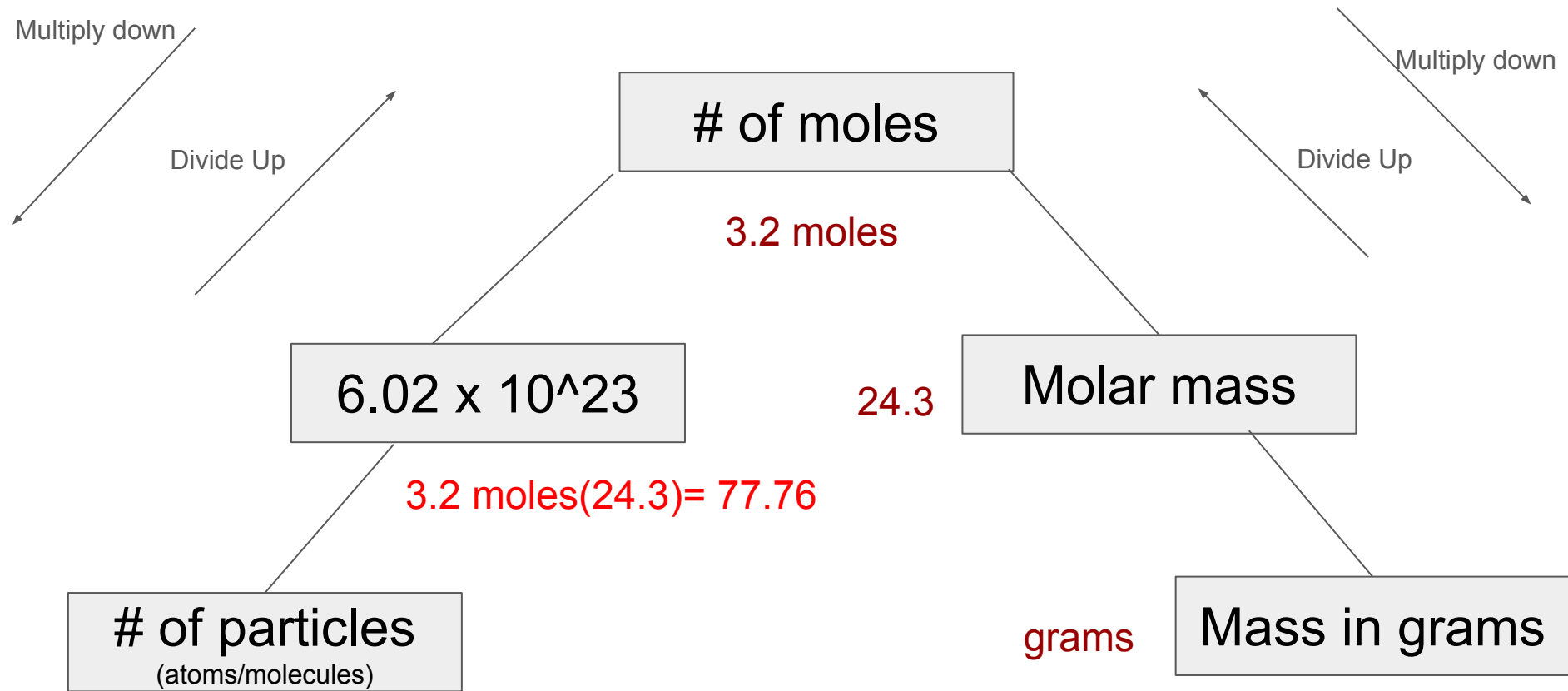
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



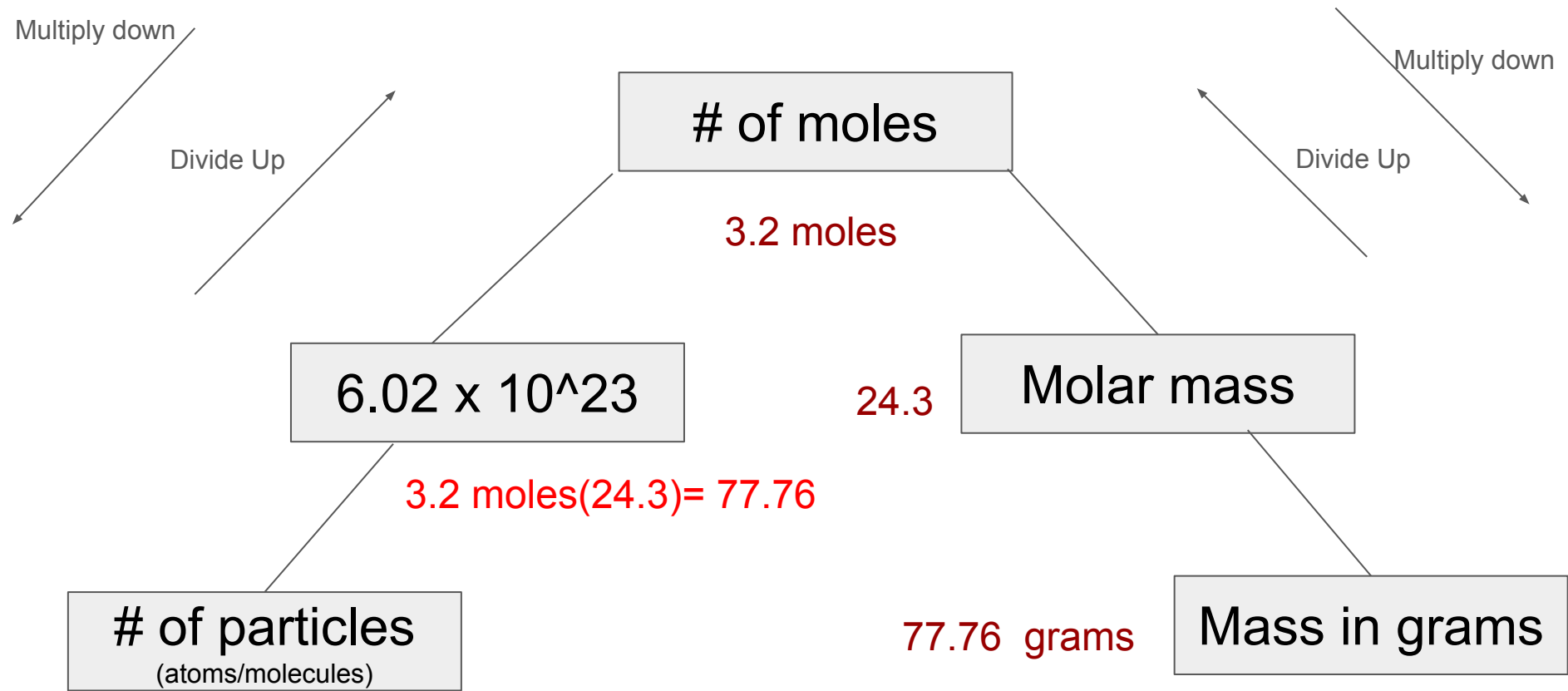
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?



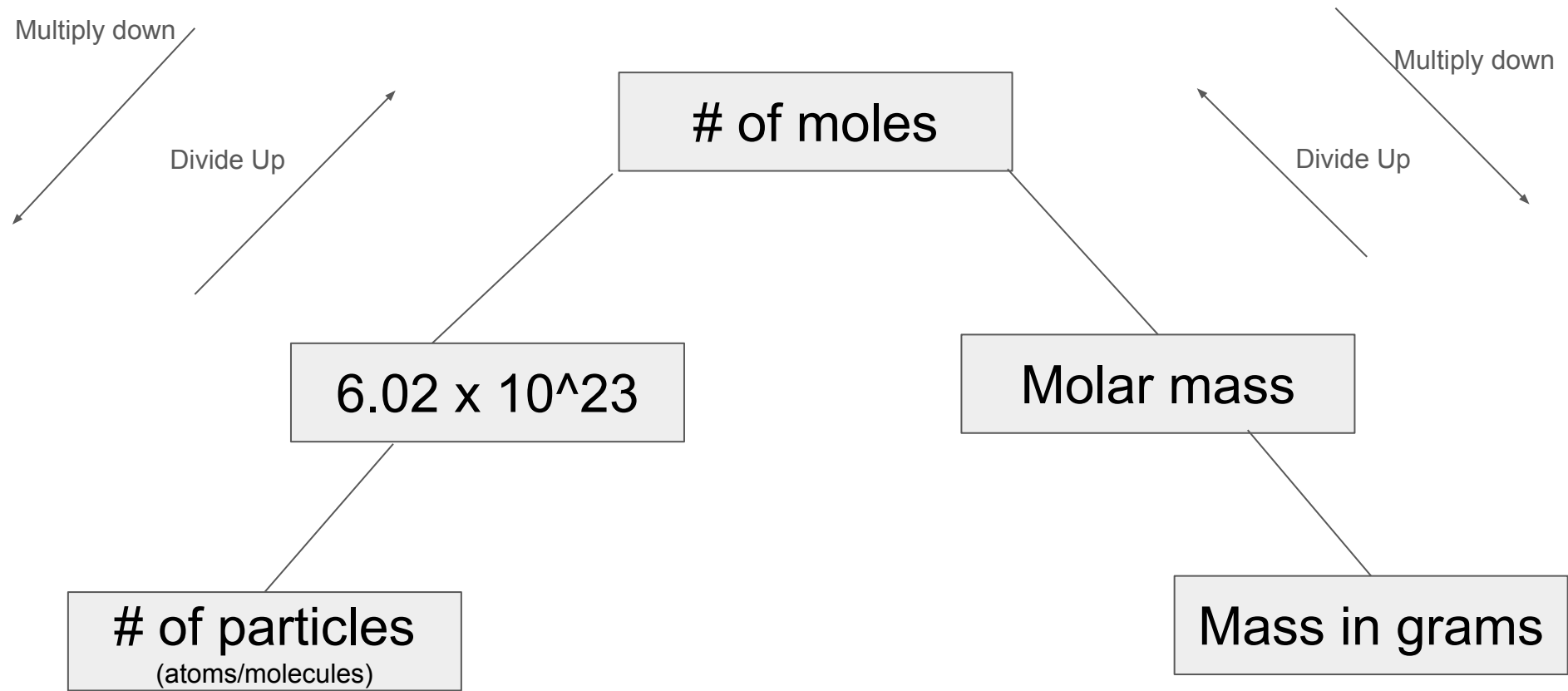
What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?

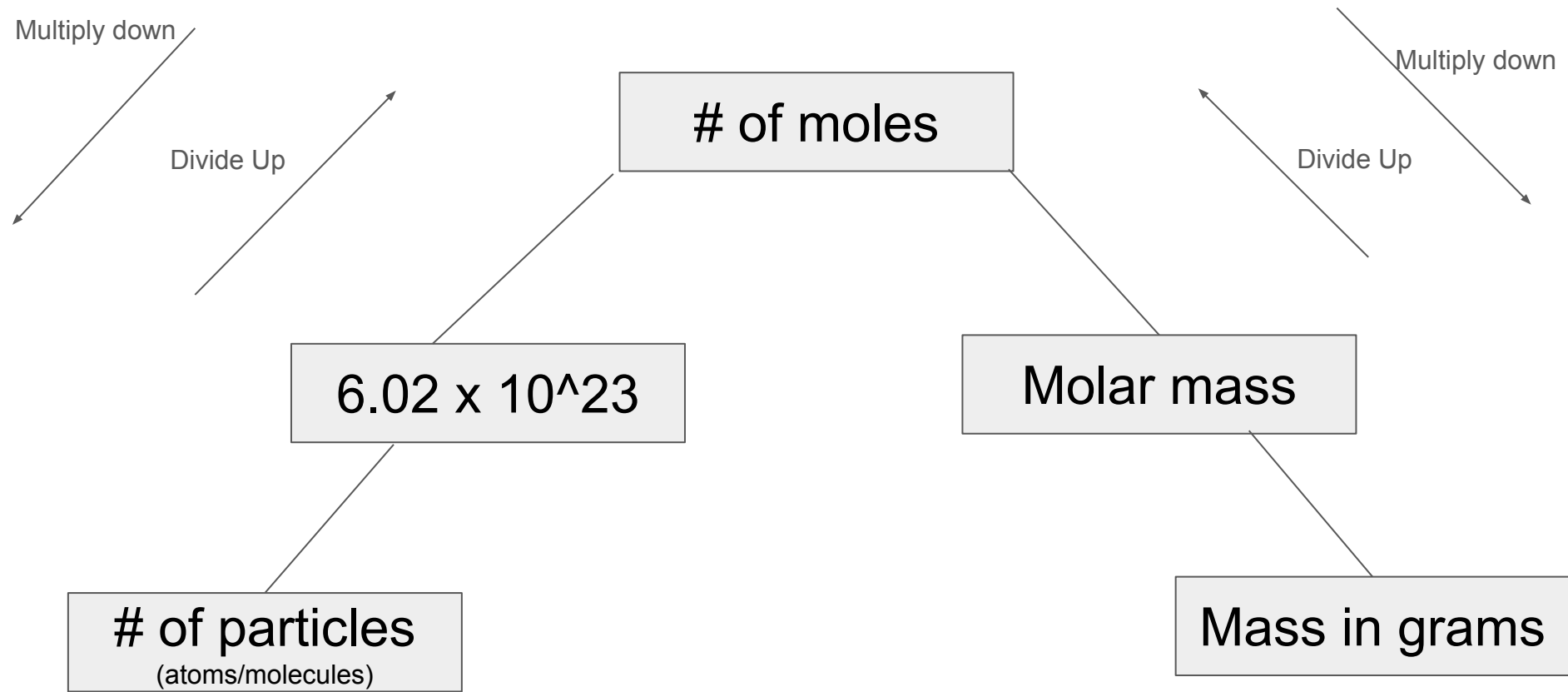


What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?

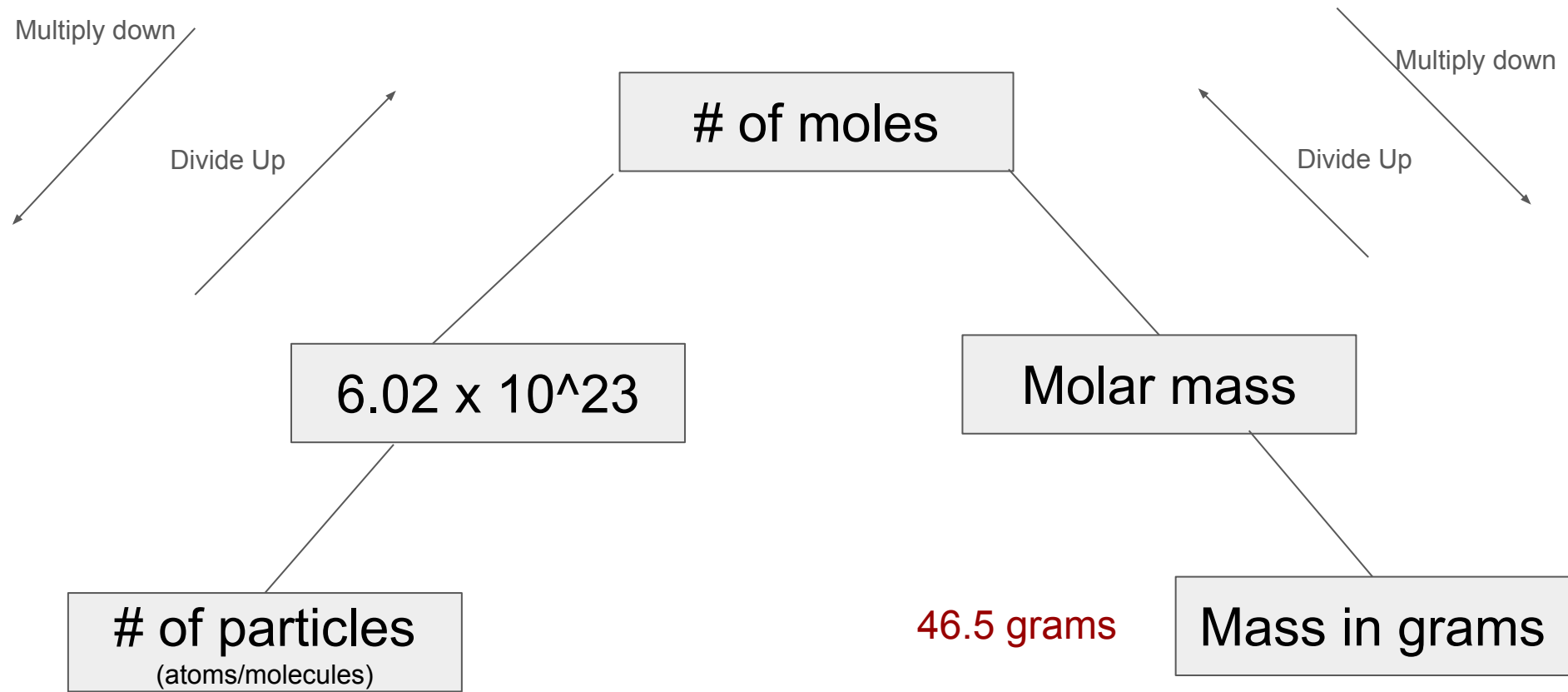


What is the mass (g) of 3.2 moles of Magnesium (Mg) atoms?

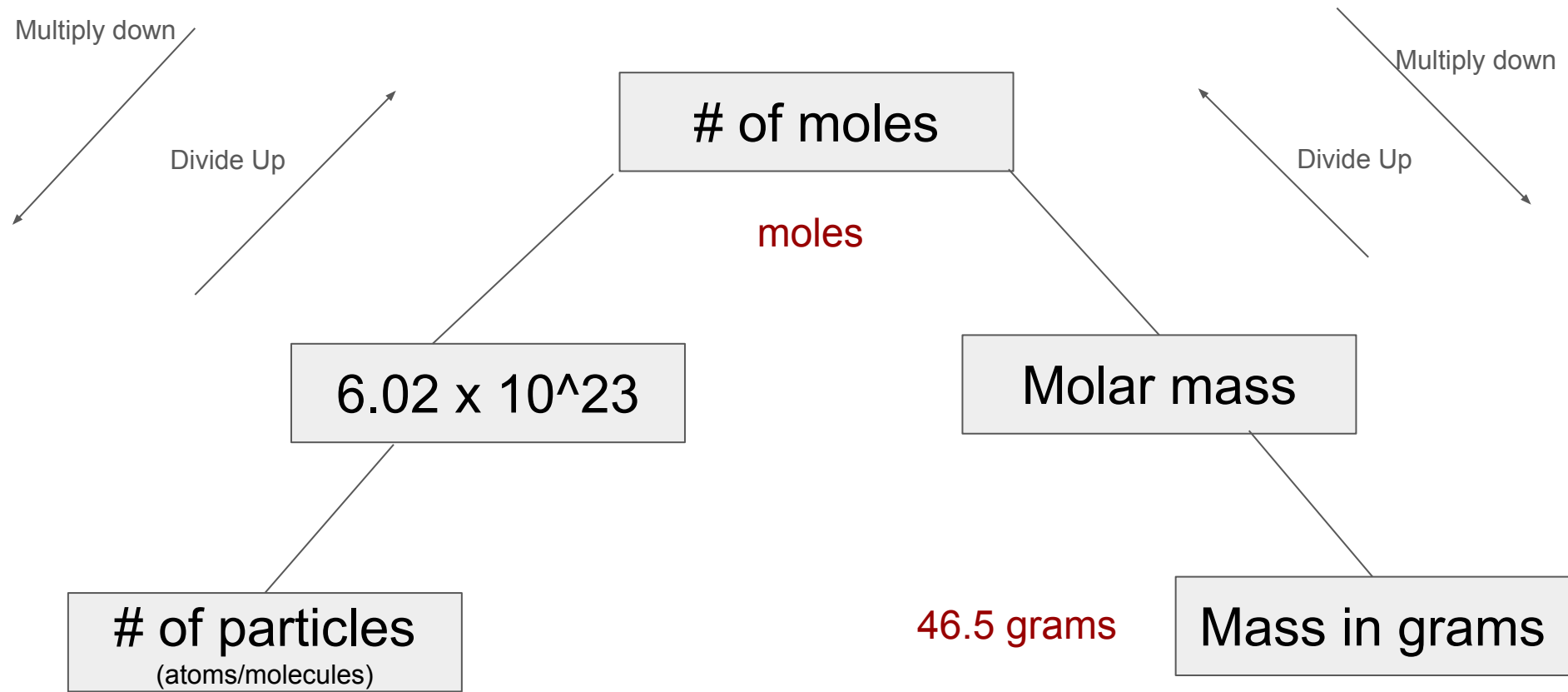




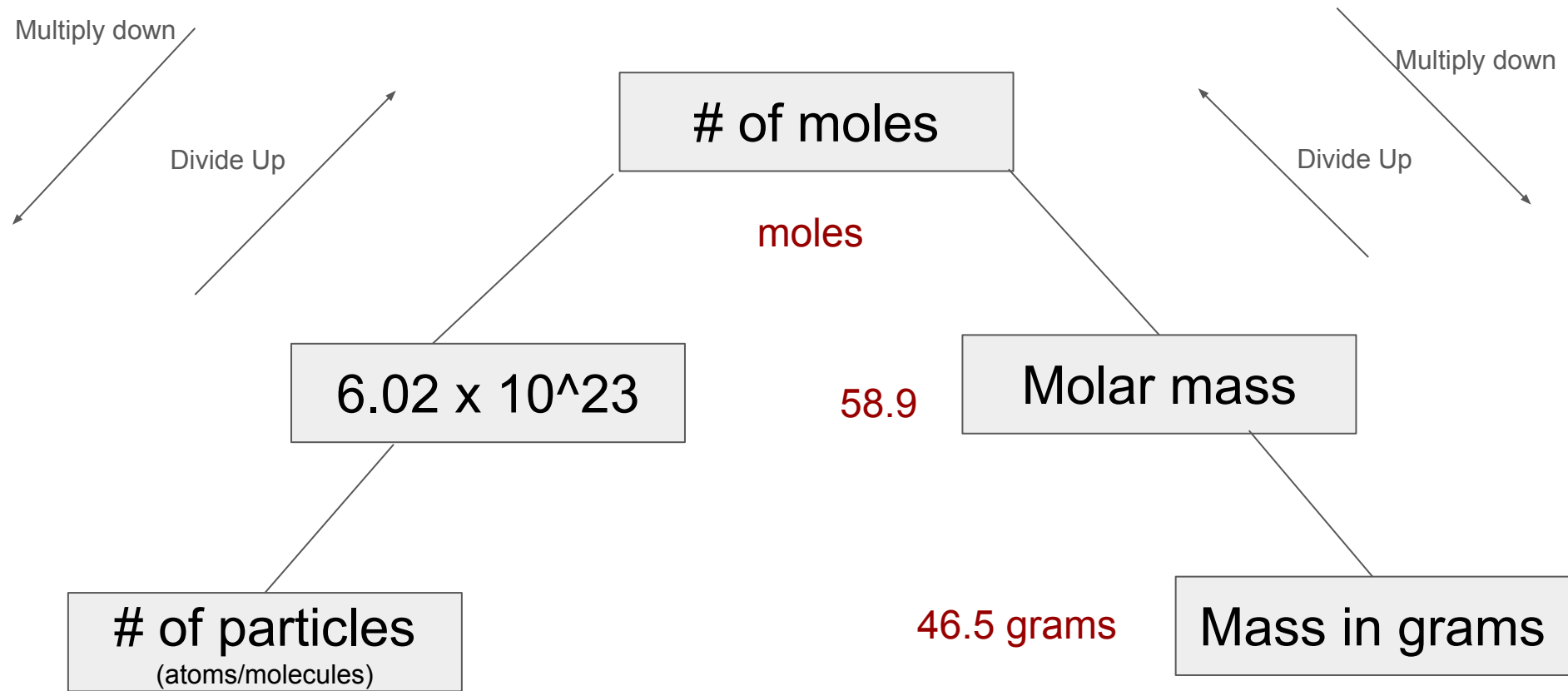
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



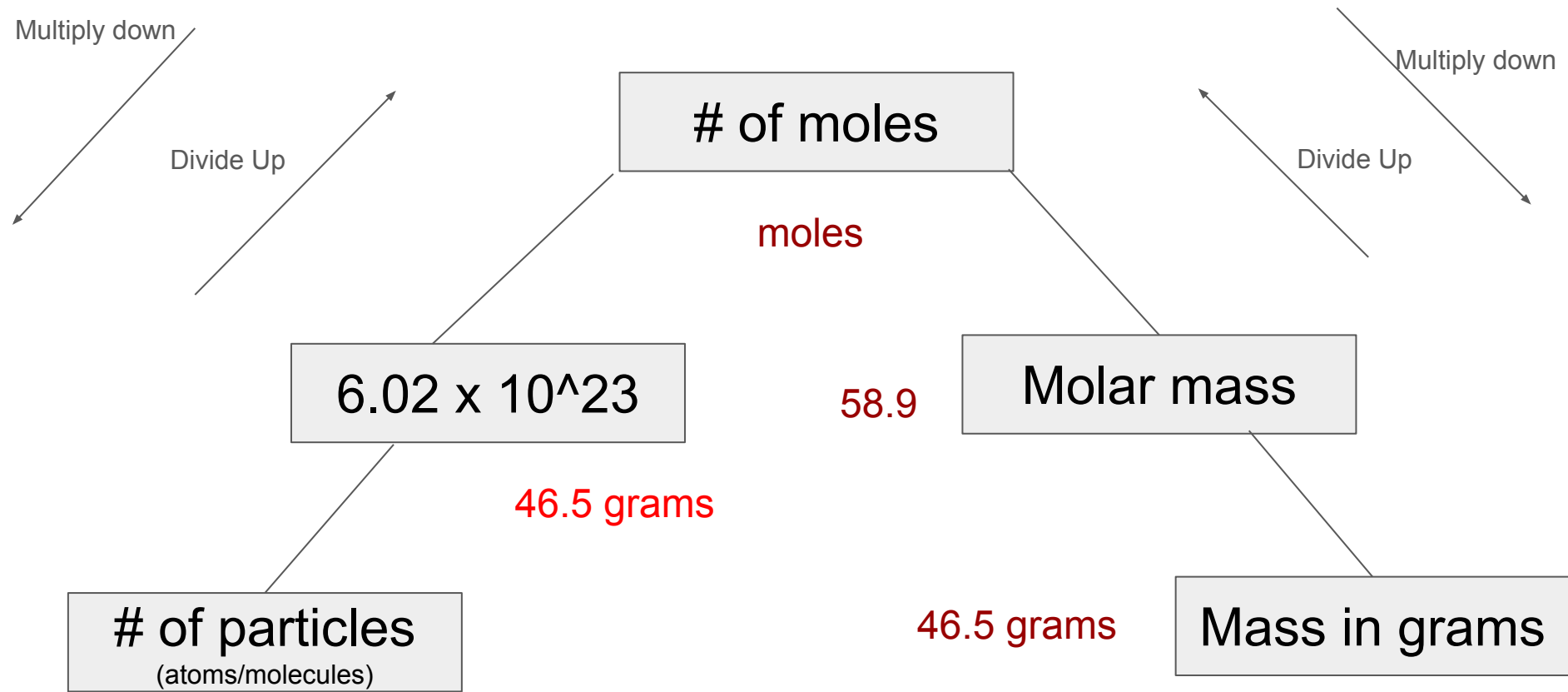
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



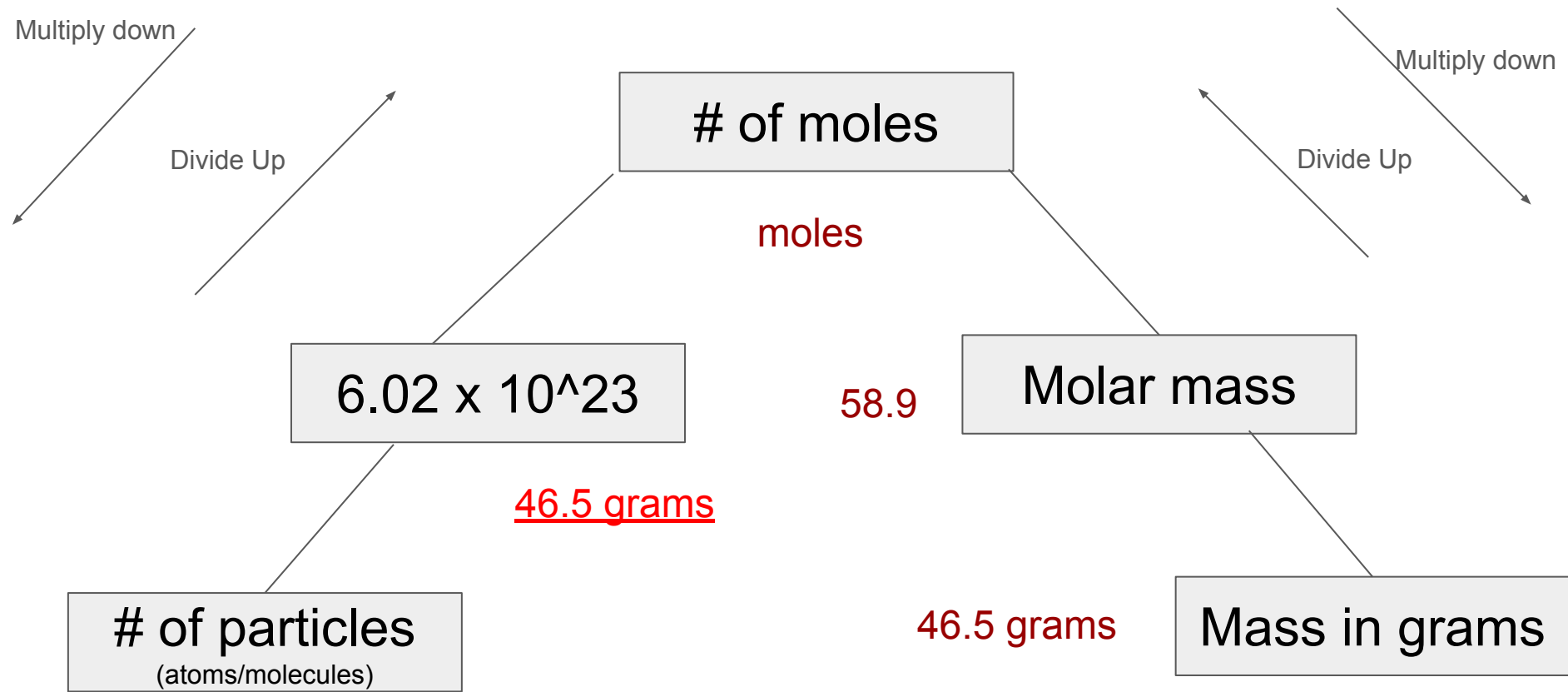
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



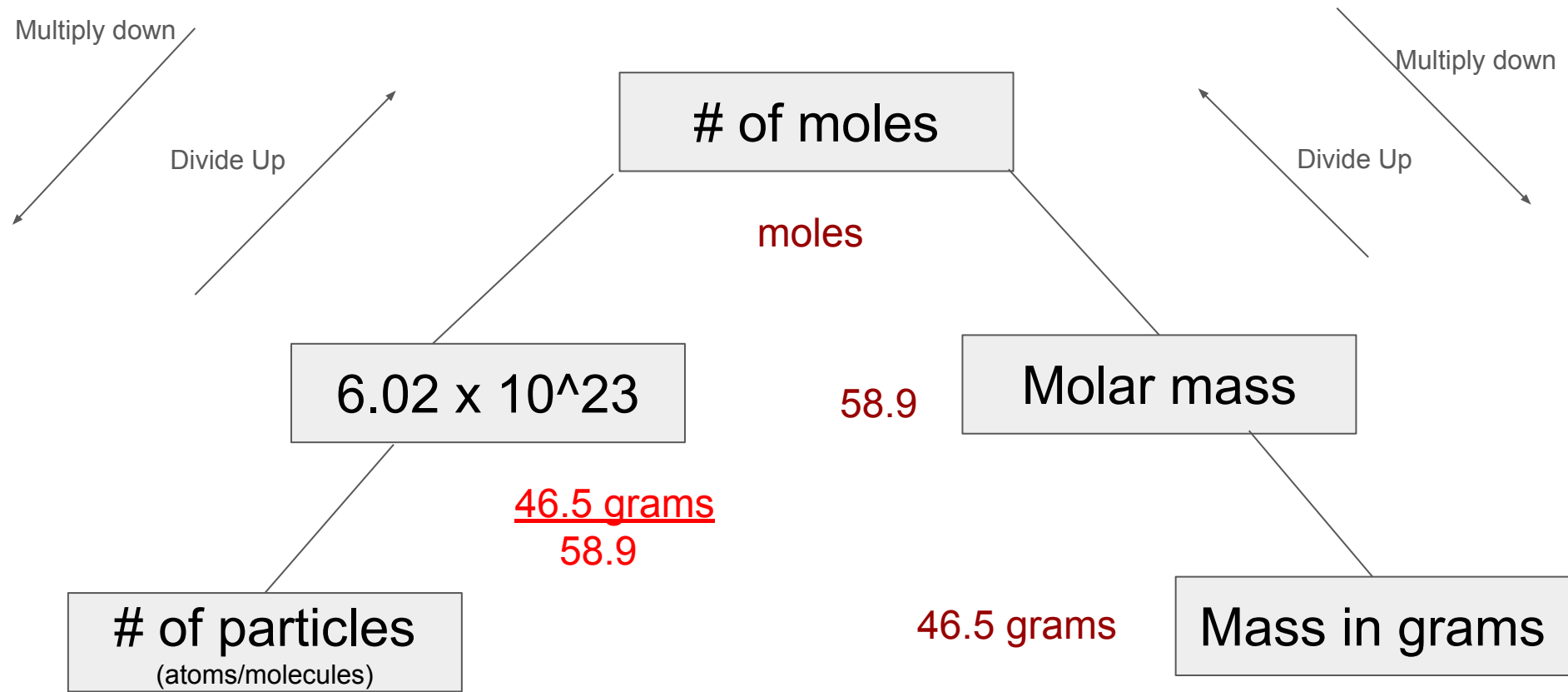
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



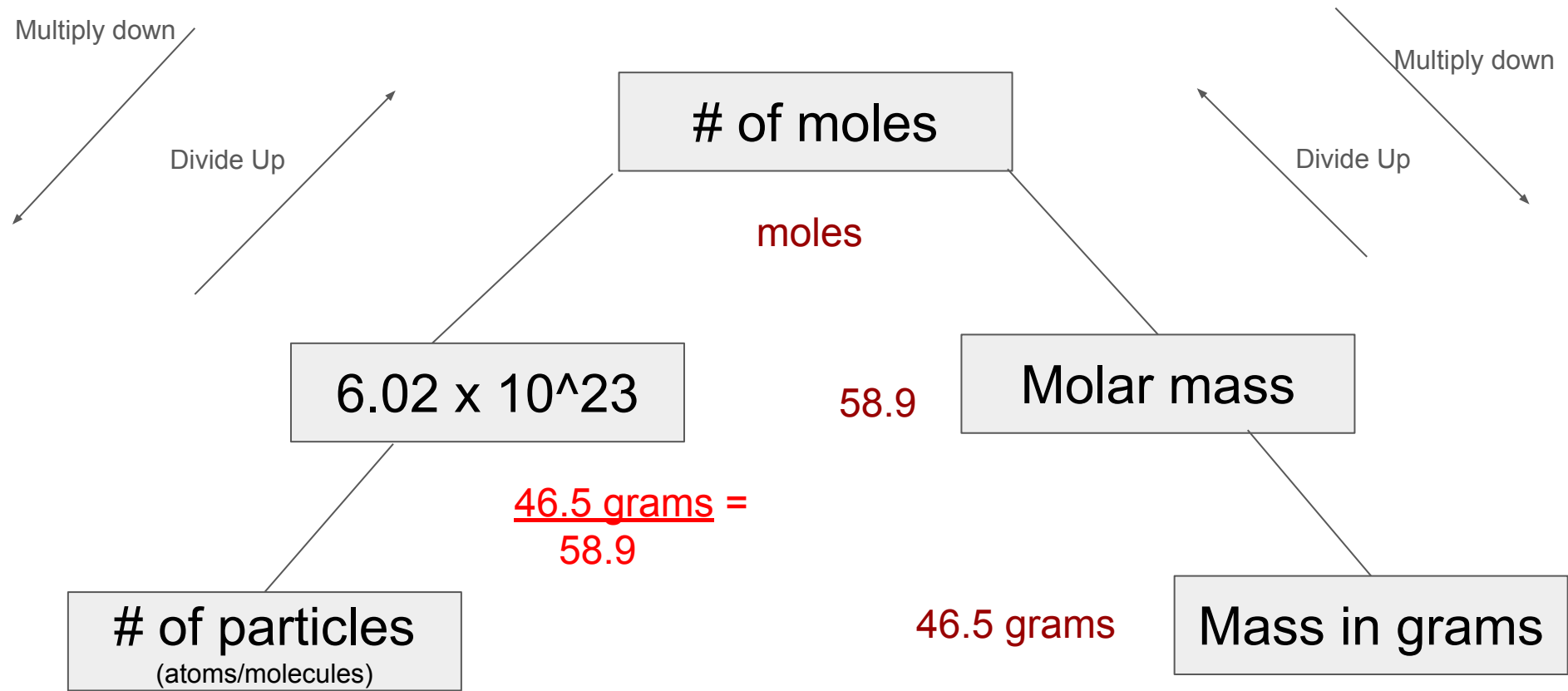
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



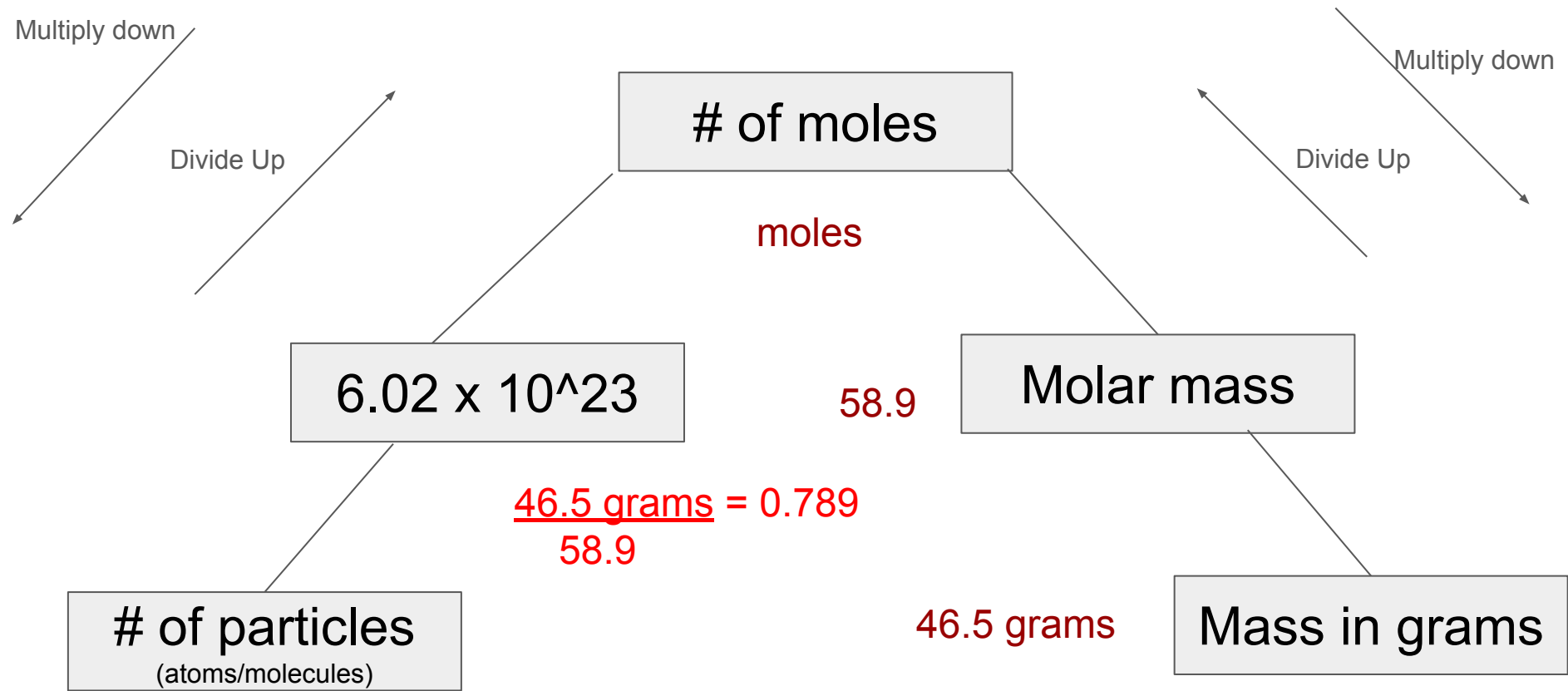
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



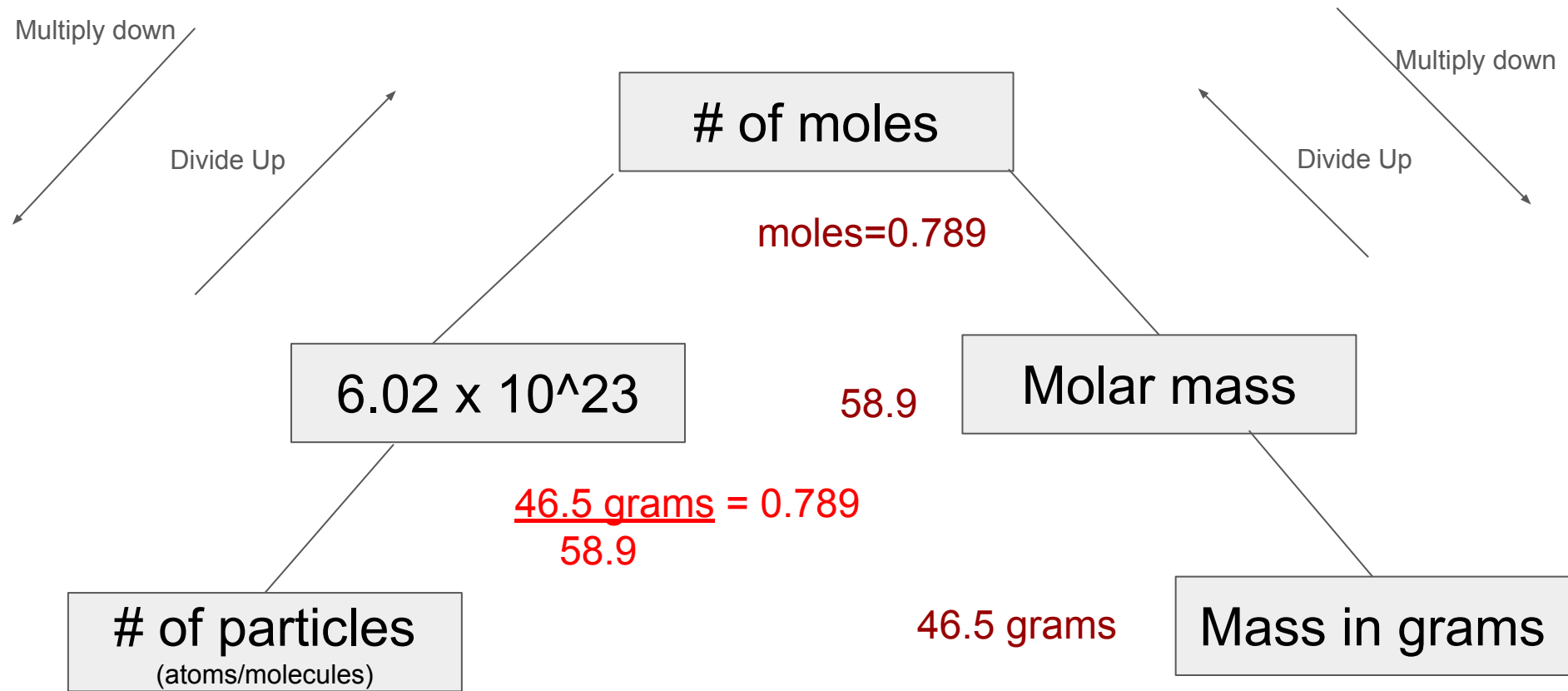
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?



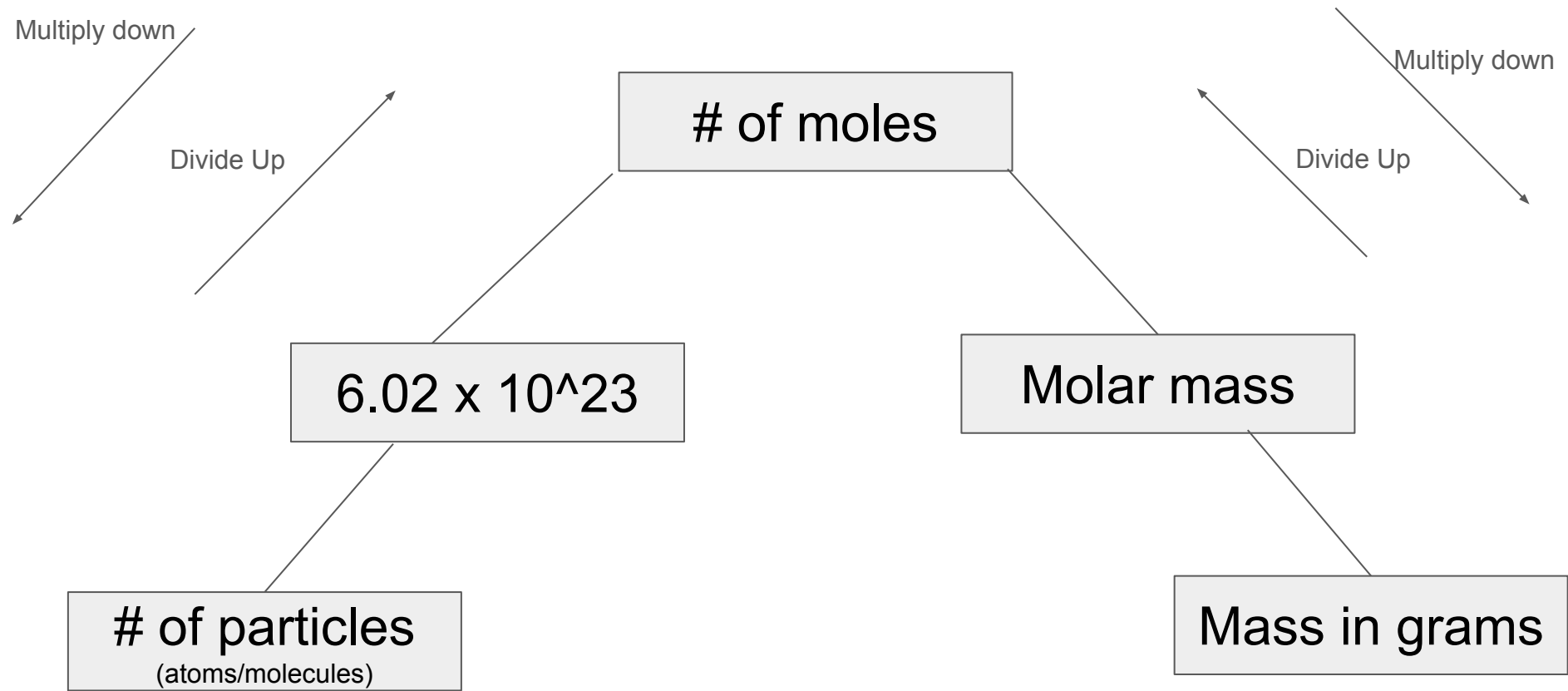
How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?

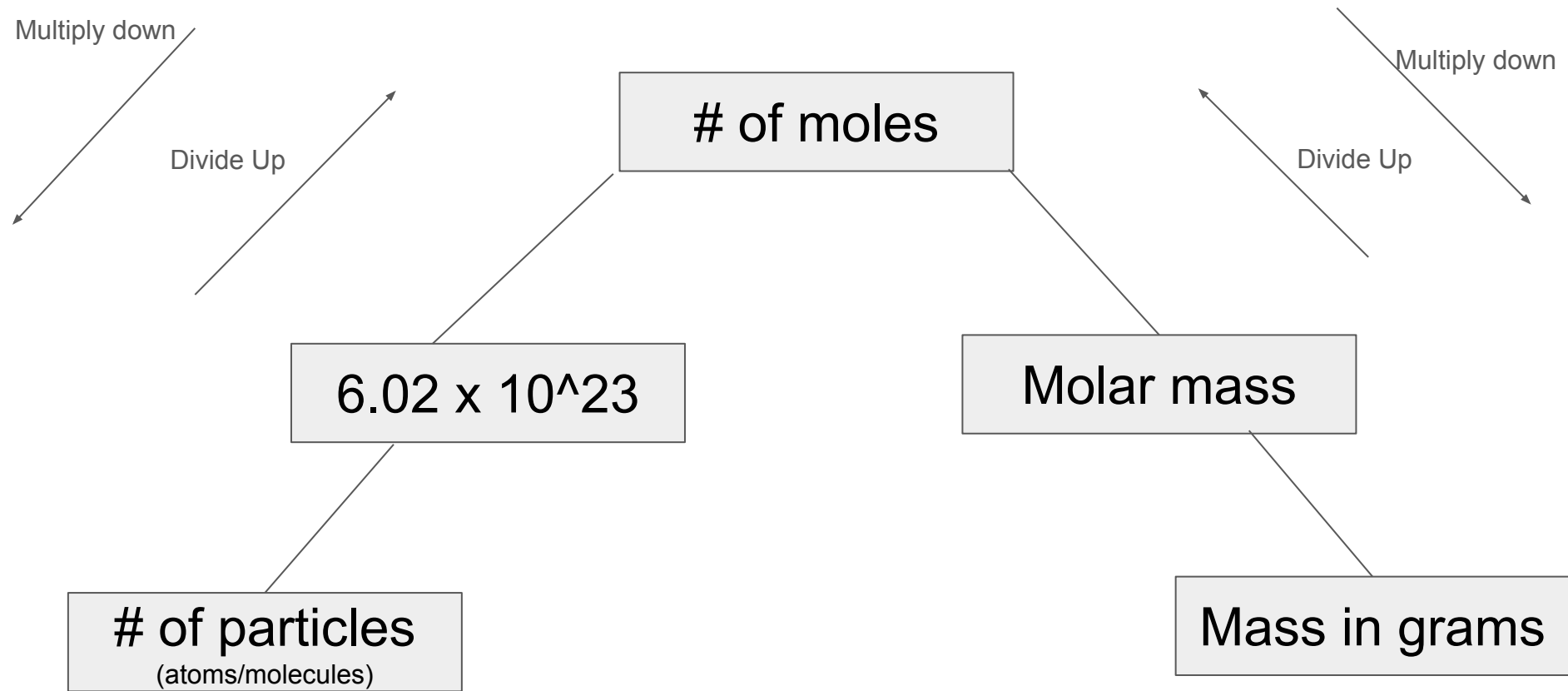


How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?

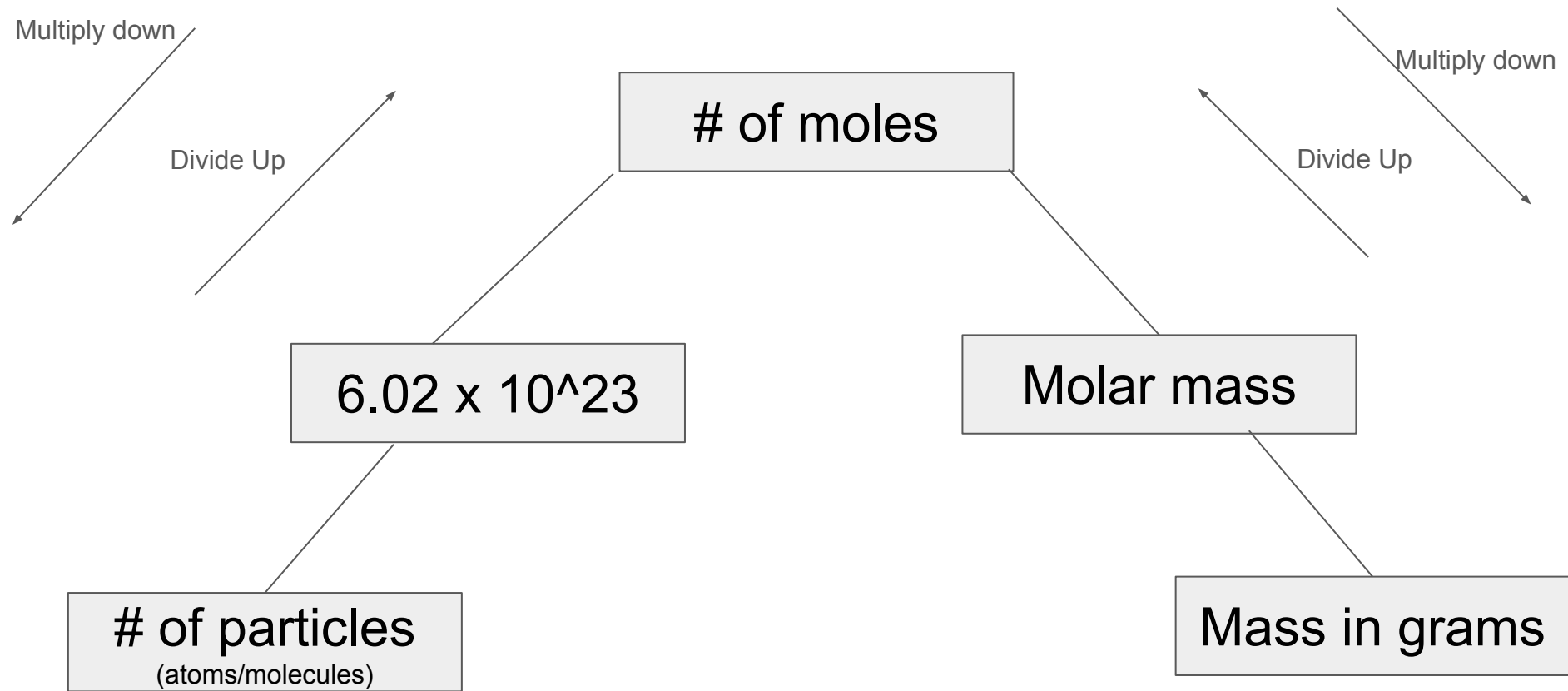


How many moles of Cobalt (Co) are in a 46.5 gram sample of Cobalt?

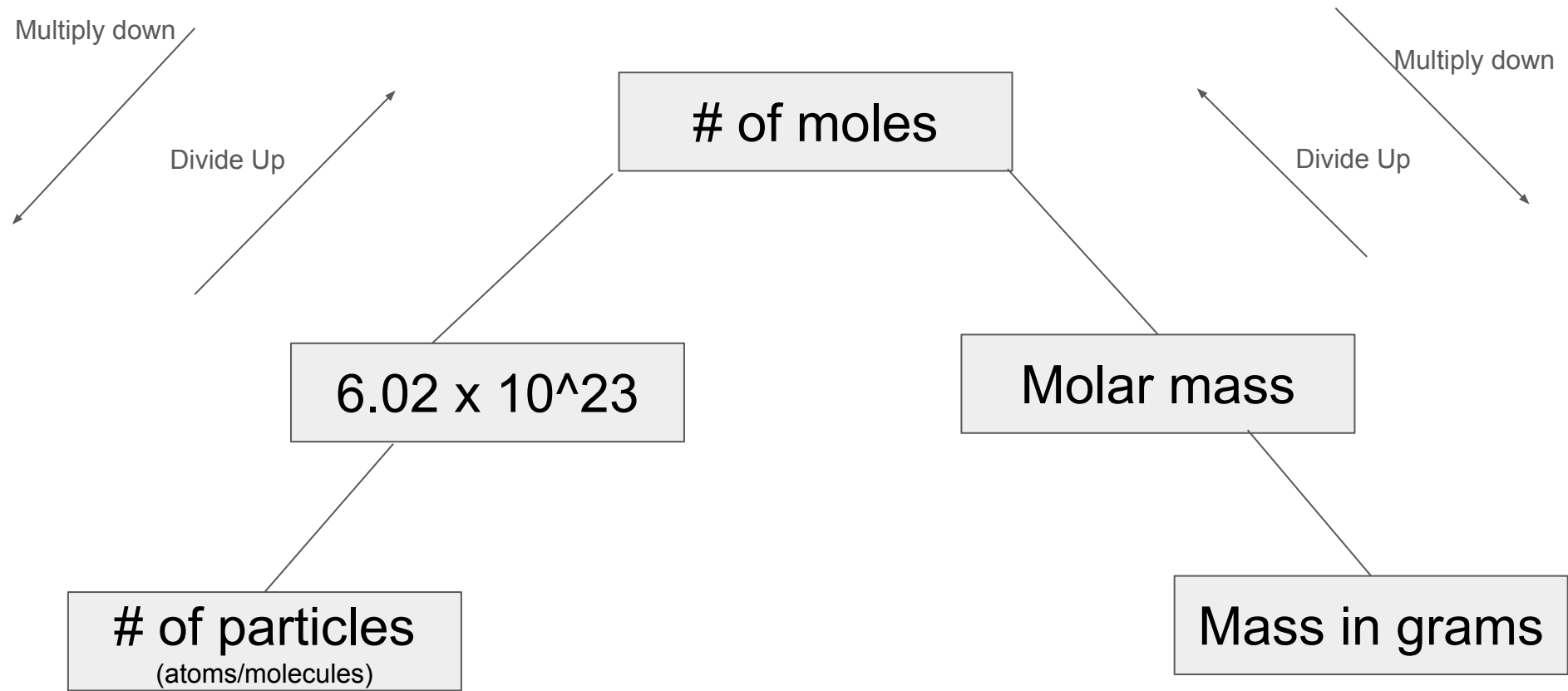




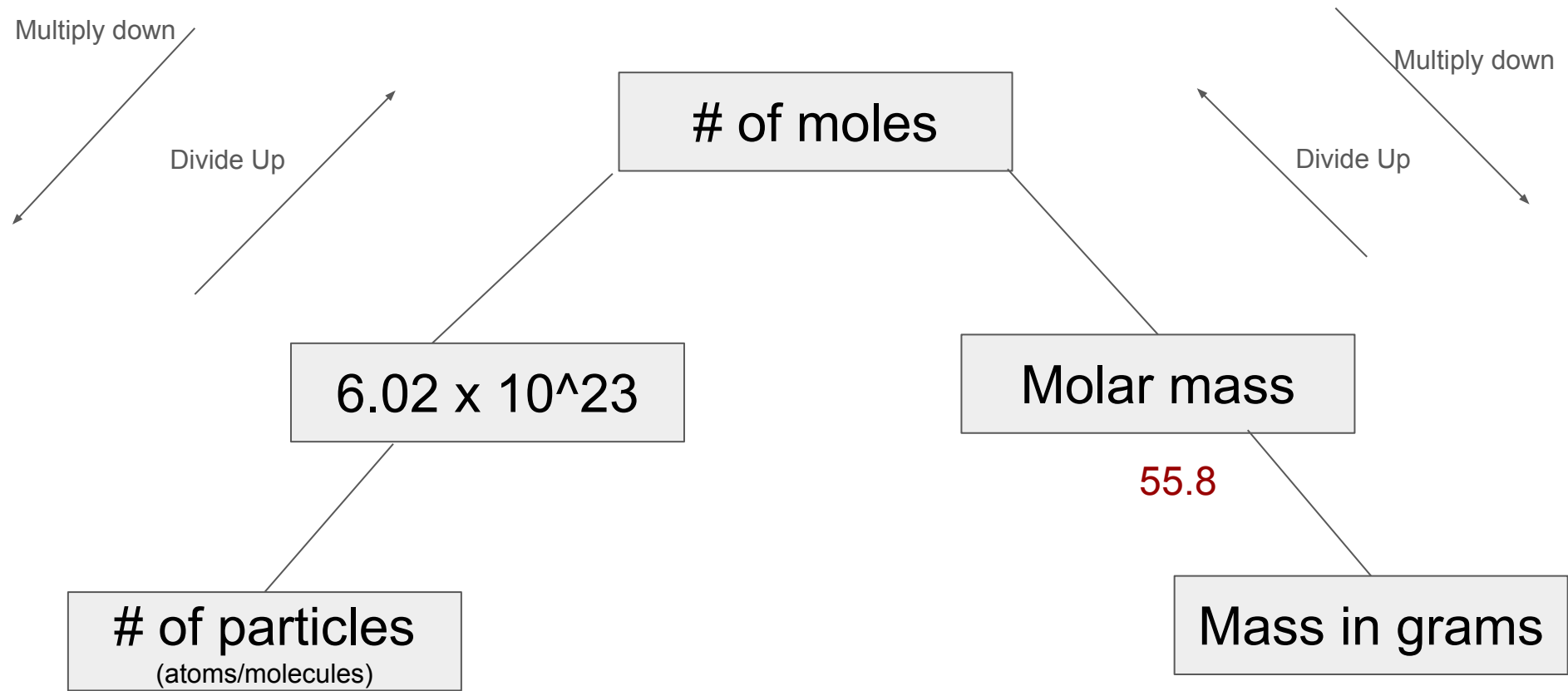
How many atoms are in a 10 gram sample of Iron?



How many atoms are in a 10 gram sample of Iron?



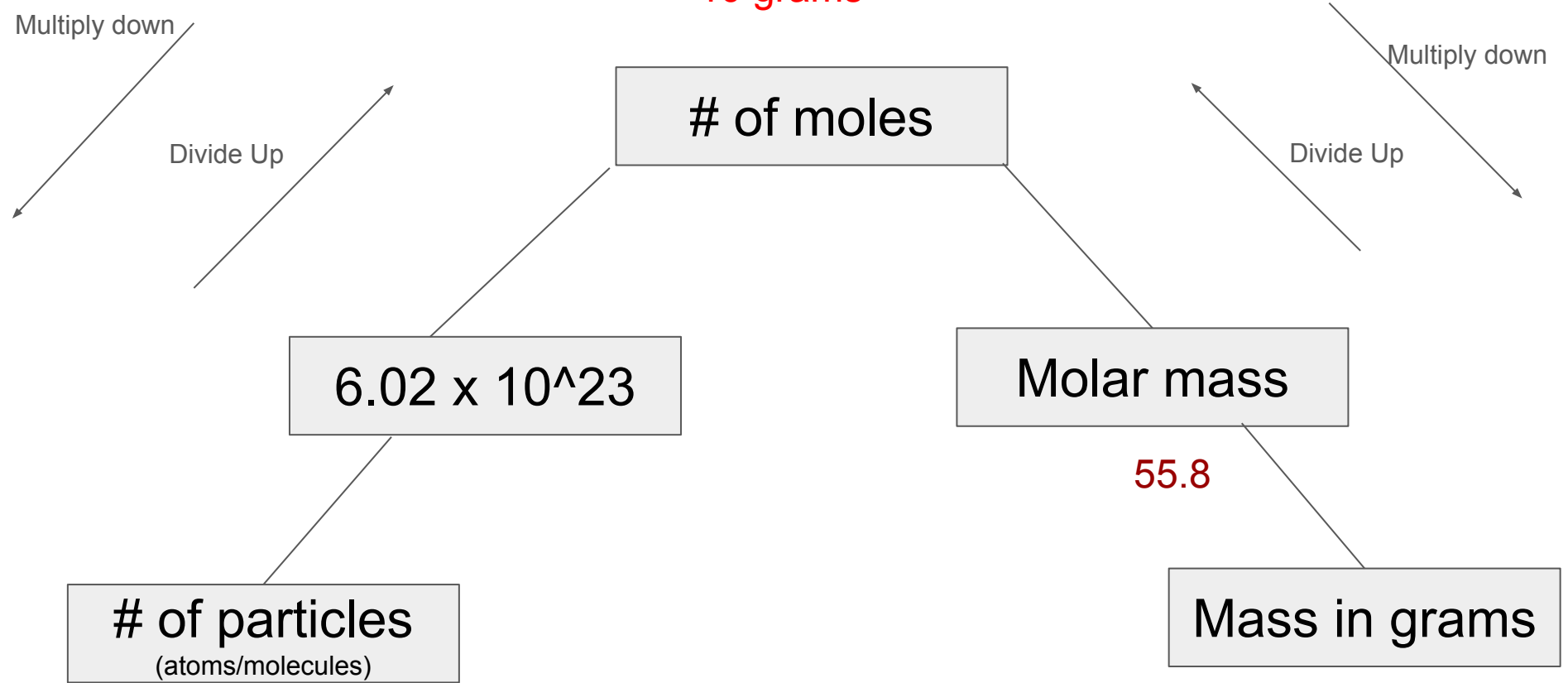
How many atoms are in a 10 gram sample of Iron? **10 grams**



How many atoms are in a 10 gram sample of Iron?

10 grams

10 grams



How many atoms are in a 10 gram sample of Iron?

10 grams

10 grams

Multiply down
Divide Up

Multiply down
Divide Up

of moles

6.02×10^{23}

Molar mass

55.8

of particles
(atoms/molecules)

Mass in grams

How many atoms are in a 10 gram sample of Iron?

10 grams

10 grams
55.8

of moles

6.02×10^{23}

Molar mass

55.8

of particles
(atoms/molecules)

Mass in grams

How many atoms are in a 10 gram sample of Iron?

10 grams

Multiply down
Divide Up

Multiply down
Divide Up

10 grams =
55.8

of moles

6.02×10^{23}

Molar mass

55.8

of particles
(atoms/molecules)

Mass in grams

How many atoms are in a 10 gram sample of Iron?

10 grams

Multiply down
Divide Up

Multiply down
Divide Up

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

6.02×10^{23}

Molar mass

55.8

of particles
(atoms/molecules)

Mass in grams

How many atoms are in a 10 gram sample of Iron?

10 grams

Multiply down
Divide Up

Multiply down
Divide Up

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

Molar mass

55.8

of particles
(atoms/molecules)

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Multiply down
Divide Up

Multiply down
Divide Up

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

0.18

of particles
(atoms/molecules)

Molar mass

55.8

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Divide Up

Divide Up

Multiply down

Multiply down

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

0.18 (6.02×10^{23})

of particles
(atoms/molecules)

Molar mass

55.8

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Divide Up

Divide Up

Multiply down

Multiply down

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

$$0.18 (6.02 \times 10^{23}) =$$

of particles
(atoms/molecules)

Molar mass

55.8

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Multiply down
Divide Up

Multiply down
Divide Up

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

$$0.18 (6.02 \times 10^{23}) = 1.08 \times 10^{23}$$

of particles
(atoms/molecules)

Molar mass

55.8

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Multiply down
Divide Up

Multiply down
Divide Up

$$\frac{10 \text{ grams}}{55.8} = 0.18$$

of moles

0.18 moles

6.02×10^{23}

$$0.18 (6.02 \times 10^{23}) = 1.08 \times 10^{23}$$

of particles
(atoms/molecules)

1.08×10^{23} atoms

Molar mass

55.8

Mass in grams

10 grams

How many atoms are in a 10 gram sample of Iron?

Multiply down
Divide Up

Multiply down
Divide Up



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